

AP Bio: Campbell Chapters 8, 9, 10 / Principles of Life Chapters 6, 24.1, 25.3, 25.4

Energy and ATP

- Energy coupling = use of exergonic process to drive endergonic process
- ATP = sugar ribose + nitrogenous base adenine + 3 phosphate groups
 - Hydrolysis breaks bonds between phosphate groups
 - Bond is broken → inorganic phosphate leaves → ATP is now ADP
 - This is an exergonic reaction
 - The bonds actually aren't exceptionally high energy, it's just that the ATP molecule has high energy compared to the products (ADP and the inorganic phosphate) so that's why ATP releases so much energy - its chemical state is being changed to a lower energy state
 - All 3 phosphate groups are negatively charged → like charges are crowded together → mutual repulsion → instability → more energy is released through hydrolysis
 - Think of the 3 phosphate groups as a compressed spring
- Cell is able to couple energy of ATP hydrolysis directly to endergonic processes by transferring phosphate group from ATP to some other molecule (which is said to be phosphorylated)
- Three types of cellular work almost always powered by ATP hydrolysis
 - Phosphate group is transferred from ATP to some other molecule → phosphorylated molecule undergoes change that performs work
- ATP is renewable, can be regenerated by adding phosphate to ADP
- Catabolic (exergonic) pathways, e.g. cellular respiration, provide energy for endergonic process of making ATP
- Changing one molecule into another generally involves contorting the starting molecule into a very unstable state before reaction can proceed
 - Reactant molecules absorb energy from their surroundings → reach contorted state where bonds can change → new bonds of product molecules form → energy is released as heat → molecules return to stable shapes with lower energy

Activation Energy

- Activation energy = initial investment of energy for starting a reaction, i.e. energy required to contort the reactant molecules so the bonds can change
 - Reactants reach a transition state, i.e. where reactants are unstable, bonds can break or form
 - Activation energy is simply the energy needed to start the reaction, it doesn't designate whether the reaction is endergonic or exergonic (both have an activation energy)

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