

AP BIOLOGY EXAM REVIEW GUIDE

"The price of success is hard work, dedication to the job at hand, and the determination that whether we win or lose, we have applied the best of ourselves to the task at hand."

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CONCEPT 1 - BIOCHEMISTRY

1. CHNOPS- most common elements in all living matter
2. Bonds- ionic (transfer electrons), covalent (sharing- polar/unequal sharing and non-polar/equal sharing), hydrogen (weak bonds between hydrogen and negatively charged items), hydrophobic interactions (how non-polar compounds congregate together- lipids)
3. pH
 - acid-base/ 0-14, # of H ions determines scale; logarithmic- $\text{pH } 3 = 10^{-3} = 1/1000$
 - blood- 7.4, stomach- 2, small intestine- 8; enzymes are specific to pH
4. Water properties- polarity, cohesion(attraction to other water molecules), adhesion (attraction to other charged compounds) low density when frozen, versatile solvent, high heat of fusion/vaporization; surface tension
5. Organic molecules - monomers are simplest form of all; monomers join together via dehydration synthesis (loss of water) to make polymers; polymers are broken down via hydrolysis (input of water)

Carbohydrates- CHO 1:2:1 ratio, monomer= monosaccharides, 2=disaccharides, 3 or more= polysaccharides

Used for energy (cell respiration)

Examples

- (1) glucose- immediate energy to make ATP
- (2) starch- stored energy in plants
- (3) glycogen- stored energy in animals (stored in liver)
- (4) cellulose- plant cell wall

Lipids – C, H, O (not a 1:2:1 ratio) *P only in phospholipids

- (1) fats, waxes, oils and sterols
- (2) Saturated fats have single bonds between carbons, unsaturated fats have at least one double bond between carbons (kinky); plants make polyunsaturated; animals make monounsaturated
- (3) Phospholipids make up cell membranes (double layer) and are amphipathic- hydrophilic and hydrophobic

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The Princeton Review



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