

# CHEMISTRY ADVANCED STUDY

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## FUNDAMENTAL CONCEPTS

### Math

- Properties of exponents
- Properties of logarithms
  - $\log_b (M^p N) = \log_b (M) + \log_b (N)$
  - $\log_b (MN) = \log_b (M) + \log_b (N)$
  - $\log_b (M^p) = p \log_b (M)$
- Graphing slope intercepts
  - $y = mx + b$
  - $b$  is the y-intercept,  $m$  is the slope
- Quadratic formula
  - $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$
  - Ex.  $x^2 + 4x - 21$
  - $x = \frac{-4 \pm \sqrt{16 - 4(1)(-21)}}{2}$
  - $x = -2 \pm 5$
  - $x = 3$  or  $x = -7$

### Science

- Scientific notation
- Dimensional analysis
- Significant figures

### Physics and Biology

- Scientific method
- Coulomb's law: electric force between like or opposite charges; dependence on separation distance and charge magnitude

## ELEMENTS AND ATOMS

Protons: define the element (atomic number), positive

Electrons: define the charge (bonds and polarity), negative

Neutrons: define isotopes, neutral

## ATOMIC MASS AND WEIGHT

Atomic Mass: Carbon-12 (exactly 12 atomic mass unit)

### Atomic Weight

- weighted average of atomic masses of element
- (atomic mass · neutrons) + (atomic mass · neutrons)

## MOLE A.K.A. AVOGADRO'S NUMBER

$$= 6.02 \cdot 10^{23}$$

### In Terms of Carbon

1 mole of  $^{12}\text{C}$  atoms in 12g of Carbon

$6.02 \cdot 10^{23} = ^{12}\text{C}$  atoms in 12g of Carbon

### Gram to AMU

1 gram = 1 mole of atomic mass units

1 gram =  $6.02 \cdot 10^{23}$  amu

1 mol  $^{27}\text{Al}$  = 27g

2 mol  $^{28}\text{Si}$  = 56g

4 mol  $^{16}\text{O}$  = 64g

# Atoms And Moles Study Guide

**Sabine Zange**



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