

Redox Exercise (Part 2)

Chapter 20 Worksheet: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} ; oxidized/reducing agent O^0 to O^{2-} ; reduced/ox. ag.
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+} ; oxidized/red. ag. S^0 to S^{2-} ; reduced/ox. ag.
- $2\text{Cs} + \text{Br}_2 \longrightarrow 2\text{CsBr}$ Cs^0 to Cs^{+} ; oxidized/red. ag. Br^0 to Br^{-} ; reduced/ox. ag.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$ Mg^0 to Mg^{2+} ; oxidized/red. ag. N^0 to N^{3-} ; reduced/ox. ag.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$ Fe^0 to Fe^{3+} ; oxidized/red. ag. O^0 to O^{2-} ; reduced/ox. ag.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$ Cl^0 to Cl^{-} ; reduced/ox. ag. Br^{-} to Br^0 ; oxidized/red. ag.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$ Si^0 to Si^{4+} ; oxidized/red. ag. F^0 to F^{-} ; reduced/ox. ag.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$ Mg^0 to Mg^{2+} ; oxidized/red. ag. H^{+} to H^0 ; reduced/o.a.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$ Na^0 to Na^{+} ; oxidized/r.a. H^{+} to H^0 ; reduced/o.a.

11. Give the oxidation number of each kind of atom or ion.

- | | | | | | | |
|------------|-------|--------------------|---------------------|---------------------|------------|-------------|
| a. sulfate | b. Sn | c. S^{2-} | d. Fe^{3+} | e. Sn^{4+} | f. nitrate | g. ammonium |
| 2- | 0 | 2- | 3+ | 4+ | 1- | 1+ |

12. Calculate the oxidation number of chromium in each of the following.

- | | | | | |
|----------------------------|---------------------------------------|--------------------|-------------|---------------|
| a. Cr_2O_3 | b. $\text{Na}_2\text{Cr}_2\text{O}_7$ | c. CrSO_4 | d. chromate | e. dichromate |
| 3+ | 6+ | 2+ | 6+ | 6+ |

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$ C^0 to C^{+4} ; oxidized S^{6+} to S^{4+} ; reduced
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$ N^{5+} to N^{2+} ; reduced I^{-} to I^0 ; oxidized
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$ Mn^{7+} to Mn^{2+} ; reduced Cl^{-} to Cl^0 ; oxidized
- $\text{Sb} + \text{HNO}_3 \longrightarrow \text{Sb}_2\text{O}_3 + \text{NO} + \text{H}_2\text{O}$ Sb^0 to Sb^{3+} ; oxidized N^{5+} to N^{2+} ; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- oxidizing agent: sulfur reducing agent: carbon
- oxidizing agent: nitrogen reducing agent: iodine
- oxidizing agent: manganese reducing agent: chloride
- oxidizing agent: nitrogen reducing agent: antimony

Balancing Redox Equations Chemistry If8766

Harold P. C. Burrell



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Principles of Redox Reactions Malika Ammam,2018-03-05 As a teacher of physical chemistry I noticed that students even in advanced classes have difficulties in understanding the basics of redox chemistry In this Section 3 I attempted to discuss some fundamental principles related to redox processes by focusing on the species that might lose or gain electrons determination of the oxidation numbers or states of atoms in compounds and ways of balancing redox reactions To further clarify the discussed concepts numerous questions and problems with detailed answers are provided Most of these questions are formulated by students like you I believe that this Section 3 would greatly help students with levels varying from high school to advanced university classes

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Balancing Redox Equations Lifelique,2019 This lesson plan covers how to balance redox equations by inspection through two systematic processes

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Balancing Organic Redox Equations Harold P. C. Burrell,1959 *Learning Channel* ,1980

Balancing Oxidation-Reduction Equations Using the Ion-Electron Method Marcia L. Gillette,H. Anthony Neidig,1995-02-01

A Review of the Methods for Balancing Oxidation-reduction Equations Henry Vivian Stewart,1962

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Redox J. Schüring,H.D. Schulz,W.R. Fischer,Jürgen Böttcher,W.H.M. Duijnsveld,2000-01-24 Few processes are as important for environmental geochemistry as the interplay

between the oxidation and reduction of dissolved and solid species The knowledge of the redox conditions is most important to predict the geochemical behaviour of a great number of components the mobilities of which are directly or indirectly controlled by redox processes The understanding of the chemical mechanisms responsible for the establishment of measurable potentials is the major key for the evaluation and sensitive interpretation of data This book is suitable for advanced undergraduates as well as for all scientists dealing with the measurement and interpretation of redox conditions in the natural environment

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