

Balancing a Chemical Equation – Rules

1. Count your atoms and set up your T-Chart.
2. You can only change coefficients, NEVER change the subscripts.
3. Balance Hydrogen last.
4. Balance Oxygen second-to-last.

Balancing a Chemical Equation – Worked Example



Note – this is a difficult problem. Most require less steps

Step 1: Count your atoms and set up your T-Chart.
Result: Nitrogen and Oxygen are balanced but Hydrogen is not balanced in the products side.



T-Chart Step 1		
Reactants	Atoms	Products
1	Nitrogen	1
2	Oxygen	2 + 1 = 3
3	Hydrogen	2

Step 2: Add more hydrogen to the products by putting a "2" coefficient in front of H_2O .
Result: the increased hydrogen in the products, now have Oxygen and Hydrogen are unbalancing, set them aside for the time being, add.



T-Chart Step 2		
Reactants	Atoms	Products
2	Nitrogen	1
2	Oxygen	2 + 2 + 2 = 6
6	Hydrogen	4 + 2 = 6

Step 3: Add more Oxygen and Oxygen to the reactants by putting a "3" coefficient in front of O_2 and O_2 .
Result: the increased Oxygen and Hydrogen in the reactants but now unbalancing in reactants, set aside Nitrogen first.



T-Chart Step 3		
Reactants	Atoms	Products
2 + 3 = 5	Nitrogen	1
6 + 6 = 12	Oxygen	2 + 2 + 2 = 6
6 + 6 = 12	Hydrogen	4 + 2 = 6

Step 4: Add more Nitrogen to the reactants by putting a "2" coefficient in front of NO .
Result: Nitrogen and Oxygen are balanced but, now Hydrogen is not balanced in the products side.



T-Chart Step 4		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2 + 2 = 4
6 + 6 = 12	Oxygen	2 + 2 + 2 + 2 = 8
6 + 6 = 12	Hydrogen	4 + 2 = 6

Step 5: Add more Hydrogen to the products by putting a "3" coefficient in front of H_2O .
Result: 3 Oxygens in reactants but are unbalanced Oxygen on the products side.



T-Chart Step 5		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2 + 2 = 4
6 + 6 = 12	Oxygen	2 + 2 + 3 + 3 = 10
6 + 6 = 12	Hydrogen	4 + 6 = 10

Balancing Chemical Equations Cheat Sheet

Heather Hattori, Richard H. Langley



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