

Stoichiometry Lab - **ANSWERS****PRE-LAB QUESTIONS:**

1. Calculate the molar mass of sodium bicarbonate (NaHCO_3). 84.003 g/mol
 2. Calculate the molar mass of acetic acid (CH_3COOH). 60.052 g/mol
 3. Calculate the molar mass of carbon dioxide (CO_2). 44.003 g/mol
 4. Vinegar is generally 4-5% acetic acid dissolved in distilled water. That means that a 5% vinegar solution contains 0.05 mol of acetic acid per liter of solution. How many moles of acetic acid are present in 30.00 mL?
- $$\frac{0.05 \text{ mol CH}_3\text{COOH}}{1 \text{ liter vinegar}} \times \frac{0.03 \text{ liter vinegar}}{1 \text{ mol vinegar}} = 0.0015 \text{ mol CH}_3\text{COOH in 30.0 mL of vinegar}$$
5. When baking soda (sodium bicarbonate) reacts with vinegar (acetic acid), they form sodium acetate, water, and carbon dioxide gas. Balance the chemical equation for this reaction.



6. A student measures out 5.00 g of baking soda to react with 50.00 mL of vinegar (which contains 0.042 moles of acetic acid).

a) Determine the limiting reagent, excess reagent, and theoretical yield (in grams) of carbon dioxide. Show your work.

$$5.00 \text{ g baking soda} \times \frac{1 \text{ mol baking soda}}{84.003 \text{ g baking soda}} = \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol baking soda}} \times \frac{5.00 \text{ g baking soda}}{1 \text{ mol carbon dioxide}} = 0.60 \text{ g carbon dioxide}$$

$$0.042 \text{ mole acetic acid} \times \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol acetic acid}} = \frac{0.042 \text{ mole acetic acid}}{1 \text{ mol carbon dioxide}} \times 44.003 \text{ g carbon dioxide} = 1.85 \text{ g carbon dioxide}$$

*** Since acetic acid produced the least amount of product, the acetic acid in vinegar is the **limiting reagent** and the baking soda is the **excess reagent**.

*** The theoretical yield of carbon dioxide is 1.85 g.

- b) This reaction will produce carbon dioxide bubbles that escape. After the reaction, the mass of the reaction decreased by 1.50 g. This is the actual mass of carbon dioxide that was produced from the reaction. Calculate the percentage yield of carbon dioxide. Show your work.

$$\text{Percent Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent Yield} = \frac{1.50 \text{ g}}{1.85 \text{ g}} \times 100 = 79.4\% \text{ yield}$$

Baking Soda And Vinegar Stoichiometry Lab Answers

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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Table of Contents Baking Soda And Vinegar Stoichiometry Lab Answers

1. Understanding the eBook Baking Soda And Vinegar Stoichiometry Lab Answers
 - The Rise of Digital Reading Baking Soda And Vinegar Stoichiometry Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Baking Soda And Vinegar Stoichiometry Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Baking Soda And Vinegar Stoichiometry Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Baking Soda And Vinegar Stoichiometry Lab Answers
 - Personalized Recommendations
 - Baking Soda And Vinegar Stoichiometry Lab Answers User Reviews and Ratings
 - Baking Soda And Vinegar Stoichiometry Lab Answers and Bestseller Lists
5. Accessing Baking Soda And Vinegar Stoichiometry Lab Answers Free and Paid eBooks
 - Baking Soda And Vinegar Stoichiometry Lab Answers Public Domain eBooks
 - Baking Soda And Vinegar Stoichiometry Lab Answers eBook Subscription Services

- Baking Soda And Vinegar Stoichiometry Lab Answers Budget-Friendly Options
- 6. Navigating Baking Soda And Vinegar Stoichiometry Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Baking Soda And Vinegar Stoichiometry Lab Answers Compatibility with Devices
 - Baking Soda And Vinegar Stoichiometry Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Baking Soda And Vinegar Stoichiometry Lab Answers
 - Highlighting and Note-Taking Baking Soda And Vinegar Stoichiometry Lab Answers
 - Interactive Elements Baking Soda And Vinegar Stoichiometry Lab Answers
- 8. Staying Engaged with Baking Soda And Vinegar Stoichiometry Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Baking Soda And Vinegar Stoichiometry Lab Answers
- 9. Balancing eBooks and Physical Books Baking Soda And Vinegar Stoichiometry Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Baking Soda And Vinegar Stoichiometry Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Baking Soda And Vinegar Stoichiometry Lab Answers
 - Setting Reading Goals Baking Soda And Vinegar Stoichiometry Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Baking Soda And Vinegar Stoichiometry Lab Answers
 - Fact-Checking eBook Content of Baking Soda And Vinegar Stoichiometry Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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