



American Water Works
Association

Errata to
AWWA Manual M11
Steel Pipe—A Guide for Design and Installation, 4th ed.
(December 2013)

Chapter 1

1. On page 1, paragraph 2, line 6 under History, the metric conversion should read 15,000 psi (103.42 MPa).
2. On page 7, paragraph 2, lines 6 and 7 under Stress and Strain, “min” should read “µin.”
3. On page 8, paragraph 2, lines 2 and 4, respectively, the metric conversions should read 30,000,000 psi (207 TPa) and 30 psi (207 kPa).
4. On page 10, paragraph 1, line 2, the metric conversion should read 30 psi (207 kPa).
5. On page 10, paragraph 1, line 4, “5,000 min/in.” should read “5,000 µin./in.”
6. On page 10, paragraph 2, line 6, “5,000 min/in.” should read “5,000 µin./in.”
7. On page 10, paragraph 2, line 8, “300,000 min/in.” should read “300,000 µin./in.”
8. On page 11, paragraph 5, lines 6 and 10, under Analysis Based on Strain, “1,750 min/in.” should read “1,750 µin./in.”
9. On page 11, paragraph 7, line 2, under Analysis Based on Strain, “5,000 min/in.” should read “5,000 µin./in.”
10. On page 18, Eq 1-1 should read
$$\sigma_n = \sigma_2 (\cos^2 \alpha) + \sigma_1 (\sin^2 \alpha) \quad (1-1)$$

Chapter 6

11. On page 62, Table 6-1, column 1, Type of Soil, replace “Coarse-grained soils with little or no fines (SP, SM, GP, GW)” with “Coarse-grained soils with little or no fines (SP, SW, GP, GW).”
12. On page 67, line 5, replace $(P = 0.7(4)(28,152) = 7,883 \text{ kg/m}^2)$ with $(P = 0.07(4)(28,152) = 7,883 \text{ kg/m}^2)$.

Awwa Manual M4

JG Myers



Awwa Manual M4:

Water Fluoridation Principles and Practices AWWA Staff,2004 *Water Fluoridation Principles and Practices* Ari Copeland,Awwa,2016-07-20 **An Introduction to Engineering of Water Supply Systems** J. Paul Guyer, P.E., R.A.,2021-04-23

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + 2\text{e}^{-} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{H}_2\text{O} + 2\text{e}^{-} \rightarrow 4\text{HCN} + 2\text{OH}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_3^- + 2\text{MnO}_2 + \text{Cr}_2\text{O}_7^{2-} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}_2\text{Te}$ 10 $\text{ReO}_4^- + 11\text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$ as $\text{NO}_3^- + \text{IO}_3^- + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{H}_2\text{NO}_2 + \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 no 2 and $\text{C}_6\text{H}_8\text{O}_6$ b no 2 and $\text{C}_6\text{H}_8\text{O}_6$ h and no 2 c h and