

- Probably not, it's water (H_2O)
- SATP** (standard ambient temperature & pressure: 25°C , 100 kPa).
STP (standard T&P: 0°C , 101.325 kPa)
Element: cannot be broken down by chemical means (all same type of atom). Compound: broken down by chemical means (two or more different types of atoms)
Group: vertical column on periodic table
Period: horizontal (left/right) row
Metal: element to the left of the staircase line.
They tend to be solids at SATP, and conduct.
Nonmetals: to the right of staircase line. Tend to be gases, non-conductors, and brittle as solid

- Thompson: with the identification of electrons he proposed that negative electrons existed in a positive dough.
Rutherford: famous gold foil experiment proved that an atom was mostly empty, with a dense positive nucleus orbited by electrons
Bohr: added to Rutherford's model the idea of "shells". Evidence includes line spectra.
- Mendeleev ordered table according to atomic mass (today it's done by atomic number)
 - Atomic number = # of protons
Mass number = # of protons + # of neutrons
They are averages (of different isotopes)
 - 20, 37, 17, 17



11. Covalent = a, d Ionic = b, c, e

12. Lewis:



- In order from low to high boiling points:
 H_2 : covalent ($\Delta\text{EN}=0$), CH_4 : covalent (0.4),
 H_2O : polar covalent (1.4), LiF : ionic (3.0)
- HCl , Na_2O , PCl_3 , Al_2O_3 , MgO
- a) copper(I) iodide, b) $\text{HI}(\text{aq})$, c) dinitrogen tetroxide, d) phosphorous acid, e) PBr_3 , f) Fe_2O_3 , g) K_2N , h) $\text{H}_2\text{C}_2\text{O}_4$, i) dichlorine heptoxide, j) hydrofluoric acid, k) nickel (II) sulfate hexahydrate, l) hydrogen sulfide
- combustion: $\text{AB} + \text{oxygen} \rightarrow \text{oxides of A \& B}$
synthesis: $\text{A} + \text{B} \rightarrow \text{C}$, decomposition: $\text{AB} \rightarrow \text{A} + \text{B}$
single displacement: $\text{A} + \text{BC} \rightarrow \text{AC} + \text{B}$
double displacement: $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$
a) S, b) DD, c) S, d) D, e) SD, f) DD, g) S, h) SD

- I - Alkali metals, II - Alkaline earth metals, VII - halogens, VIII - noble gases, middle - transition elements/metals, bottom - inner transition elements (lanthanides, actinides)
- Democritus: first to propose atom and the void based purely on logic
Aristotle: his ideas were also based on thought. He was way off, but his ideas persisted for 2000 years
Dalton: 1st to investigate structure of matter by experiment. His five postulates included the idea that all matter is made of atoms, each element has its own type of atoms, and atoms are rearranged in chemical reactions

- Size increases down a group (more shells), it decreases left to right as the # of protons increases, pulling outer electrons closer
- Ionization energy: energy required to remove outer electron. It is high when atoms are small (high in group) with lots of protons (right in period). Electron affinity: the energy change when an electron adds to an atom. It is also high when atoms are small (high in group) with lots of protons (right in period).
- Electronegativity: ability of atoms, when bonded, to attract electrons (essentially a numerical value for electron affinity). It follows the same trend as electron affinity for the same reason



- Ionic: high melting/boiling points, soluble in polar solvents, conducts when dissolved in water but not as solid, brittle.
Covalent: low melting/boiling points, soluble in non-polar solvents, doesn't conduct, soft. These differences are caused by the different strength of intermolecular forces in ionic versus covalent molecules.

- a) $\text{Ca} + \text{CuSO}_4 \rightarrow \text{Cu} + \text{CaSO}_4$
b) $\text{FeCl}_2 + \text{Ag} \rightarrow \text{NR}$
c) $\text{H}_2\text{O} + \text{Ca} \rightarrow \text{H}_2 + \text{Ca}(\text{OH})_2$
d) $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2 + \text{Al}_2(\text{SO}_4)_3$
e) $\text{Na} + \text{Ni}_3(\text{PO}_4)_2 \rightarrow \text{Ni} + \text{Na}_3\text{PO}_4$
f) $\text{Au} + \text{HCl} \rightarrow \text{NR}$

Bonding Of Atoms Chemistry Study Guide Answers

Hongru Du



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