

# Structure, Bonding and properties



## 1. Why do atoms bond together?

Some atoms are very reluctant to combine with other atoms and exist in the air around us as single atoms. These are the **Noble Gases** and have **very stable electron arrangements** eg **2, 2.8** and **2.8.8** because **their outer shells are full**. The first three are shown in the diagrams below and explains why Noble Gases are so reluctant to form compounds with other elements.



**All other atoms therefore, bond together to become electronically more stable**, that is to become like Noble Gases in electron arrangement. Bonding produces new substances and usually involves only the 'outer shell' or 'valency' electrons

The phrase **CHEMICAL BOND** refers to the strong electrical force of attraction between the atoms or ions in the structure. The combining power of an atom is sometimes referred to as its **valency** and its value is linked to the number of outer electrons of the original uncombined atom.

**COVALENT BONDING** - **sharing electrons** to form **molecules** with **covalent bonds**, the bond is usually formed between **two non-metallic elements** in a molecule.

**IONIC BONDING** - By one atom **transferring electrons** to another atom.

**COORDINATE COVALENT BONDING** A coordinate covalent bond is special because it involves a shared pair of electrons that came from a single atom.

**METALLIC BONDING** The **crystal lattice of metals** consists of **ions NOT atoms** surrounded by a 'sea of electrons' forming **giant lattice**. These free or 'delocalised' electrons are the 'electronic glue' holding the particles together. There is a **strong electrical force of attraction** between these **mobile electrons (-)** and the 'immobile' **positive metal ions (+)** and this is the **metallic bond**.

- An ion is an atom or group of atoms carrying an overall **positive** or **negative** charge
  - eg  $\text{Na}^+$ ,  $\text{Cl}^-$ ,  $[\text{Cu}(\text{H}_2\text{O})]^{2+}$ ,  $\text{SO}_4^{2-}$  etc.
- If a particle, as in a neutral atom, has equal numbers of protons (+) and electrons (-) the particle charge is zero ie no overall electric charge.
- The proton number in an atom does not change BUT the number of associated electrons can!
- If negative electrons are removed the excess charge from the protons produces an overall positive ion.
- If negative electrons are gained, there is an excess of negative charge, so a negative ion is formed.
- The charge on the ion is numerically related to the number of electrons transferred.
- For any atom or group of atoms, for every electron gained you get a one unit increase in negative charge, for every electron lost you get a one unit increase in the positive

The atom **losing electrons** forms a **positive ion (cation)** and is usually a **metal**.

The atom **gaining electrons** forms a **negative ion (anion)** and is usually a **non-metallic element**.

# Atoms And Bonding Assessment Study Guide

**James Adams, Maryn Munley, Nathan  
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