

$$\begin{aligned}
 B' &= \frac{B}{RT} = \frac{-(110 \text{ cm}^3 \text{ mol}^{-1})(10^{-3} \text{ L cm}^{-3})}{(0.08314 \text{ L bar K}^{-1} \text{ mol}^{-1})(400 \text{ K})} \\
 &= -3.31 \times 10^{-3} \text{ bar}^{-1} \\
 Z &= 1 + B'P = 1 - (3.31 \times 10^{-3} \text{ bar}^{-1})(50 \text{ bar}) \\
 &= 0.835
 \end{aligned}$$

7. 1.5 Show that the virial equation written in terms of pressure has the following form for a van der Waals gas $Z = 1 + [b - (a/RT)](P/RT)$ if terms in P^2 , P^3 , etc., are neglected.

SOLUTION

$$P = \frac{RT}{V-b} - \frac{a}{V^2} \qquad \frac{PV}{RT} = \frac{V}{V-b} - \frac{a}{VRT}$$

Adding and subtracting $\frac{V-b}{V-b}$ on the right side of the equation

$$\frac{PV}{RT} = 1 + \frac{b}{V-b} - \frac{a}{VRT}$$

Since the second and third terms are small correction terms, b can be ignored in comparison with V ; and V in the second and third term can be replaced by RT/P .

$$Z = 1 + \frac{bP}{RT} - \frac{aP}{(RT)^2} = 1 + \left[b - \frac{a}{RT} \right] \left[\frac{P}{RT} \right]$$

- 1.6 Show that for a gas of spherical molecules b in the van der Waals equation is four times the molecular volume times Avogadro's constant.

SOLUTION

The molecular volume for a spherical molecule is

$$\frac{4}{3}\pi\left(\frac{d}{2}\right)^3 = \frac{\pi}{6}d^3$$

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Robert A. Alberty



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