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CHAPTER 5 : Electrons in Atoms

Section 2 : Quantum Theory and the Atom

Bohr's Model of the Atom

Energy states of hydrogen

He describes electrons in terms of their **energy levels**:

- Electrons can be only certain distances from the **nucleus**.
- Each distance corresponds to a certain quantity of **energy** that an electron can have.
- An electron that is as close to the nucleus as it can is in its **lowest** energy level, the farther an electron is from the nucleus, the **higher** the energy level that the electron occupies.
- The difference in energy between two energy levels is known as **Quantum** of energy.
- An electron can be in only one energy level or another **not** between energy levels.
- Normally if an electron is in a state of lowest possible energy, it is in a **ground state**.
- If an electron gains **energy**, it moves to an **excited state**.
- An electron in an excited state will release a specific quantity of **energy** as it quickly **falls** back to its ground state.
- The smaller the electron's orbit, the **lower** the atom's energy state, or energy level.
- Conversely, the larger the electron's orbit, the **higher** the atom's energy state, or energy level.
- Thus, a hydrogen atom can have many different **excited** states, although it contains only one electron.
- Bohr's idea is illustrated in figure 1.9

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