

AP Biology Study Guide Unit 1

Properties of Water:

1. Polarity - allows H₂O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
 - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
 - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
 - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
 - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
 - Form Stable Covalent Bonds
 - Can form double and triple bonds
 - Versatile

Biological Transition Metals - Mn, Fe, Co, Cu, Zn

- Can exist in multiple oxidation states → can be both oxidized and reduced

Biological Ions: Na⁺, Mg²⁺, Cl⁻, K⁺, Ca²⁺

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

5 Foundation Biological Molecules:

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**Neil A. Campbell, Jane B. Reece, Fred
W. Holtzclaw, Theresa Knapp Holtzclaw**

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