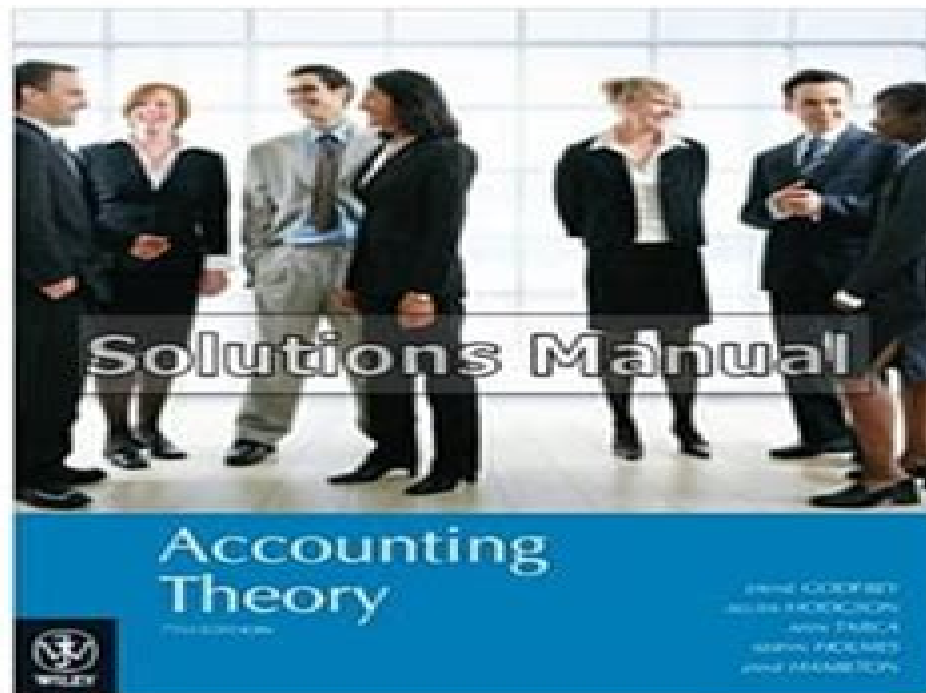


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 $\text{mg}$  s magnesium  $2 \text{hno}_3$  l nitric acid  $\text{mg}$  no  $3_2$  s magnesium nitrate h  $2$  g hydrogen gas oxygen and hydrogen get  
 automatically balanced by balancing nitrogen

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 g  $\text{h}_2\text{o}$  l b  $\text{mg}^{2+}$  aq  $\text{co}_3^{2-}$  aq  $2\text{h}^+$  aq  $2\text{no}_3^-$  aq  $\text{mg}$  write the balanced equation when nitric acid reacts with calcium hydroxide to  
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 reaction the  $\text{mg}$  replaces the  $\text{h}$  in the  $\text{hno}_3$

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 each compound to balance  $\text{mg}$   $\text{hno}_3$   $\text{mg}$  no  $3_2$   $\text{no}_2$   $\text{h}_2\text{o}$  you ll need to be

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web magnesium hydroxide nitric acid balanced total equation  $\text{mg}$  oh  $2$  aq  $2 \text{hno}_3$  aq  $2 \text{h}_2\text{o}$   $\text{mg}$  no  $3_2$  aq total ionic equation  
 $\text{mg}^{2+}$   $2 \text{oh}^-$   $2 \text{h}^+$   $2 \text{no}_3^-$   $2 \text{h}_2\text{o}$   $\text{mg}^{2+}$   $2 \text{no}_3^-$  net ionic equation  $2 \text{oh}^-$   $2 \text{h}^+$   $2 \text{h}_2\text{o}$  sulfuric acid and lithium hydroxide  
 balanced total equation  $\text{h}_2\text{so}_4$  aq  $2 \text{lioh}$  aq  $2 \text{h}_2\text{o}$