

Chapter 2 Active Reading Guide

The Chemical Context of Life

This chapter covers the basics that you may have learned in your chemistry class. Whether your teacher goes over this chapter, or assigns it for you to review on your own, the questions that follow should help you focus on the most important points.

Section 1

1. Define and give an example of the following terms:

Term	Definition
matter:	All physical substances.
element:	Chemical substances that cannot be broken down
compound:	A mix of two or more elements.

2. What four elements make up 96% of all living matter?

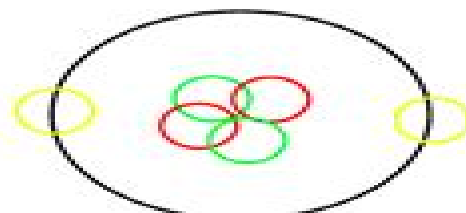
Carbon, Oxygen, Hydrogen, and Nitrogen

3. What is the difference between an *essential element* and a *trace element*?

While essential elements are needed in large quantities for an organism to live a healthy life, trace elements are needed in very little amounts.

Section 2

4. Sketch a model of an atom of helium, showing the electrons, protons, and neutrons, and atomic nucleus.



5. What is the atomic number of helium? 2 Its atomic mass? 4

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M Walker



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potassium nitrate has an orthorhombic crystal structure at room temperature which transforms to a trigonal system at 128 c 262 f on cooling from 200 c 392 f another trigonal phase forms between 124 c 255 f and 100 c 212 f sodium nitrate is isomorphous with calcite the most stable form of calcium carbonate whereas room temperature potassium nitrate is isomorphous with aragonite a slightly less stable polymorph of

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