

11.1

Given: $x = 4t^3 - 6t^2 + 2t - 1$ $x = m$, $t = s$
 Find: x , v , and a at $t = 2s$

Have - $x = 4t^3 - 6t^2 + 2t - 1$
 Then $v = \frac{dx}{dt} = 12t^2 - 12t + 2$
 And $a = \frac{dv}{dt} = 24t - 12$

At $t = 2s$: $x_2 = 4(2)^3 - 6(2)^2 + 2(2) - 1$ OR $x_2 = 19m$
 $v_2 = 12(2)^2 - 12(2) + 2$ OR $v_2 = 38 \frac{m}{s}$
 $a_2 = 48(2) - 12(2)$ OR $a_2 = 60 \frac{m}{s^2}$

11.2

Given: $x = 3t^4 + 4t^3 - 7t^2 - 5t + 8$ $x = mm$, $t = s$
 Find: x , v , and a at $t = 3s$

Have - $x = 3t^4 + 4t^3 - 7t^2 - 5t + 8$
 Then $v = \frac{dx}{dt} = 12t^3 + 12t^2 - 14t - 5$
 And $a = \frac{dv}{dt} = 36t^2 + 24t - 14$

At $t = 3s$: $x_3 = 3(3)^4 + 4(3)^3 - 7(3)^2 - 5(3) + 8$ OR $x_3 = 121mm$
 $v_3 = 12(3)^3 + 12(3)^2 - 14(3) - 5$ OR $v_3 = 305 \frac{mm}{s}$
 $a_3 = 36(3)^2 + 24(3) - 14$ OR $a_3 = 382 \frac{mm}{s^2}$

11.3

Given: $x = 6t^3 - 8 + 40 \cos 6t$ $x = m$, $t = s$
 Find: x , v , and a at $t = 6s$

Have - $x = 6t^3 - 8 + 40 \cos 6t$
 Then $v = \frac{dx}{dt} = 18t^2 - 40 \sin 6t$
 And $a = \frac{dv}{dt} = 36t - 40 \pi^2 \cos 6t$

At $t = 6s$: $x_6 = 6(6)^3 - 8 + 40 \cos 36$ OR $x_6 = 248m$
 $v_6 = 12(6)^2 - 40 \sin 36$ OR $v_6 = 71 \frac{m}{s}$
 $a_6 = 72 - 40 \pi^2 \cos 36$ OR $a_6 = 383 \frac{m}{s^2}$

11.4

Given: $x = \frac{5}{2}t^3 - \frac{7}{2}t^2 - 30t + 8$ $x = ft$, $t = s$
 Find: t , x , and a when $v = 0$

Have - $x = \frac{5}{2}t^3 - \frac{7}{2}t^2 - 30t + 8$
 Then $v = \frac{dx}{dt} = 5t^2 - 7t - 30$
 And $a = \frac{dv}{dt} = 10t - 7$

When $v = 0$: $5t^2 - 7t - 30 = 5(t^2 - t - 6) = 0$
 OR $t = 3s$ AND $t = -2s$ (reject) $\therefore t = 3s$

At $t = 3s$: $x_3 = \frac{5}{2}(3)^3 - \frac{7}{2}(3)^2 - 30(3) + 8$ OR $x_3 = -58.5ft$
 $a_3 = 10(3) - 7$ OR $a_3 = 23 \frac{ft}{s^2}$

11.5

Given: $x = 6t^4 - 2t^3 - 12t^2 + 3t + 3$ $x = m$, $t = s$
 Find: t , x , and a when $a = 0$

Have - $x = 6t^4 - 2t^3 - 12t^2 + 3t + 3$
 Then $v = \frac{dx}{dt} = 24t^3 - 6t^2 - 24t + 3$
 And $a = \frac{dv}{dt} = 72t^2 - 12t - 24$

When $a = 0$: $72t^2 - 12t - 24 = 12(6t^2 - t - 2) = 0$
 OR $(3t + 2)(2t - 1) = 0$

OR $t = -\frac{2}{3}s$ AND $t = \frac{1}{2}s$ (reject) $\therefore t = \frac{1}{2}s$
 At $t = \frac{1}{2}s$: $x_{\frac{1}{2}} = 6(\frac{1}{2})^4 - 2(\frac{1}{2})^3 - 12(\frac{1}{2})^2 + 3(\frac{1}{2}) + 3$ OR $x_{\frac{1}{2}} = 2.59m$
 $a_{\frac{1}{2}} = 72(\frac{1}{2})^2 - 12(\frac{1}{2}) - 24$ OR $a_{\frac{1}{2}} = -9.5 \frac{m}{s^2}$

11.6

Given: $x = 3t^3 - 4t^2 - 12t + 5$ $x = m$, $t = s$
 Find: (a) t when $v = 0$
 (b) x and total distance traveled when $t = 4s$

Have - $x = 3t^3 - 4t^2 - 12t + 5$
 Then $v = \frac{dx}{dt} = 9t^2 - 8t - 12$
 And $a = \frac{dv}{dt} = 18t - 8$

(a) When $v = 0$: $9t^2 - 8t - 12 = 3(3t^2 - 4t - 4) = 0$
 OR $(3t + 2)(t - 2) = 0$

OR $t = -\frac{2}{3}s$ AND $t = 2s$ (reject) $\therefore t = 2s$

(b) At $t = 4s$: $x_4 = 3(4)^3 - 4(4)^2 - 12(4) + 5$ OR $x_4 = 55m$
 $a = 18(4) - 8$ OR $a = 60 \frac{m}{s^2}$

First observe that $0 \leq t \leq 2s$: $v < 0$
 $t > 2s$: $v > 0$

Now - At $t = 0$: $x_0 = 5m$
 At $t = 2s$: $x_2 = 3(2)^3 - 4(2)^2 - 12(2) + 5 = -19m$



Then $x_2 - x_0 = -19 - 5 = -24m$

$x_4 - x_2 = 55 - (-19) = 74m$

$\therefore$ Total distance traveled = $(24 + 74)m = 98m$

11.7

Given: $x = t^3 - 9t^2 + 24t - 8$ $x = m$, $t = s$
 Find: (a) t when $v = 0$
 (b) x and total distance traveled when $a = 0$

Have - $x = t^3 - 9t^2 + 24t - 8$
 Then $v = \frac{dx}{dt} = 3t^2 - 18t + 24$
 And $a = \frac{dv}{dt} = 6t - 18$

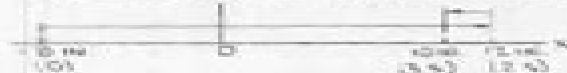
(a) When $v = 0$: $3t^2 - 18t + 24 = 3(t^2 - 6t + 8) = 0$
 OR $(t - 2)(t - 4) = 0$

OR $t = 2s$ AND $t = 4s$

(b) When $a = 0$: $6t - 18 = 0$ OR $t = 3s$
 At $t = 3s$: $x_3 = 3(3)^3 - 9(3)^2 + 24(3) - 8$ OR $x_3 = 10m$

First observe that $0 \leq t \leq 2s$: $v > 0$
 $2s < t < 4s$: $v < 0$

Now - At $t = 0$: $x_0 = -8m$
 At $t = 2s$: $x_2 = (2)^3 - 9(2)^2 + 24(2) - 8 = 12m$



Then $x_2 - x_0 = 12 - (-8) = 20m$

$x_3 - x_2 = 10 - 12 = -2m$

$\therefore$ Total distance traveled = $(20 + 2)m = 22m$

Beer Dynamics 9th Edition Solution Manual

Sabine Zange



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web how many protons neutrons and electrons are in atoms of these isotopes write the complete electron configuration for each isotope answer co has 27 protons 27 electrons and 33 neutrons $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$ i has 53 protons 53 electrons and 78 neutrons $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^5$

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web learning objectives relate the electron configurations of the elements to the shape of the periodic table determine the expected electron configuration of an element by its place on the periodic table previously we introduced the periodic table as a tool for organizing the known chemical elements

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web the answer is rather simple if you understand electron configurations the shape of the periodic table mimics the filling of the subshells with electrons let us start with h and he their electron configurations are $1s^1$ and $1s^2$ respectively with he the n 1 shell is filled

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web photoelectron spectroscopy electron configurations google classroom you might need periodic table using s p d f notation what is the electron configuration for a neutral atom of beryllium choose 1 answer $1s^2 2s^2$ a $1s^2 2s^2 1s^2 2s^1$ 2 p 1 b $1s^2 2s^1 2p^1$ 1 s 2 2 p 2 c $1s^2 2p^2$ 1 s 2 1 p 2 d $1s^2 1p^2$

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web jun 14 2015 this list of electron configurations of elements contains all the elements in increasing order of atomic number to save room the configurations are in noble gas shorthand this means part of the electron configuration has been replaced with the element symbol of the noble gas symbol

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web electron configurations and trends answer these discussion questions given the following part of an electron configuration 3s² explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund s rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is 1s 2 2s 2 2p 6 3s 2 3p 6 4s 1 or simply put ar 4s 1 q2 what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the

periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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web electron configurations are a simple way of writing down the locations of all of the electrons in an atom as we know the positively charged protons in the nucleus of an atom tend to attract negatively charged electrons

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