

# Equilibrium Practice Questions

Name: \_\_\_\_\_

## Equilibrium Questions

1. At an unspecified temperature a 3.00 L flask contains 3.5 mol of  $\text{HI(g)}$ , 4.1 mol of  $\text{H}_2\text{(g)}$  and 0.30 mol of  $\text{I}_2\text{(g)}$ . Calculate  $K_c$  for:  $\text{H}_2\text{(g)} + \text{I}_2\text{(g)} \rightleftharpoons 2\text{HI(g)}$
2. At 1200 °C,  $K_c = 2.51 \times 10^4$  (no units) for:  $\text{H}_2\text{(g)} + \text{I}_2\text{(g)} \rightleftharpoons 2\text{HI(g)}$ . What concentration of  $\text{HI(g)}$  will be produced if 0.250 mol/L of each reactant comes to equilibrium?
3. At 460°C  $K_c$  for:  $\text{SO}_2\text{(g)} + \text{NO}_2\text{(g)} \rightleftharpoons \text{SO}_3\text{(g)} + \text{NO(g)}$  - is 85.0 (no units). Estimate the product equilibrium concentrations if 0.100 mol/L of each reactant is allowed to come to equilibrium?
4. At 1530 °C, the  $K_c$  of:  $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightleftharpoons 2\text{NO(g)}$  - is  $1.20 \times 10^{-4}$  (no units). What are the equilibrium concentrations of all the reaction components, after 1.00 mol/L of each of the reactants is allowed to come to equilibrium?

# Aqueous Equilibrium Practice Problems

**Thomas Bond, Chris Hughes**



## **Aqueous Equilibrium Practice Problems:**

*Working Method Approach for Introductory Physical Chemistry Calculations* Brian J Hathaway, Clair Murphy, Brian Murphy, 2019-05-16 A Working Method Approach for Introductory Physical Chemistry Calculations is a concise inexpensive introduction to first year chemistry that is aimed at students who are weak in chemistry or have no chemistry on entry to university Such students usually find physical chemistry the most difficult part of the chemistry course and within this section numerical problem solving is an additional difficulty The text should also be invaluable to first year intending chemists This text provides an introduction to physical chemistry and the gas laws followed by chapters on thermodynamics chemical equilibrium electrochemistry and chemical kinetics Each section involves a brief introduction followed by a representative examination question which is broken down into a proposed working method Both short multiple choice questions and related full examination type questions are included This book will prove invaluable to students who need encouragement in a logical approach to problem solving in physical chemistry teaching them to think for themselves when faced with a problem

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*A Problem-Solving Approach to Aquatic Chemistry* James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including

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**Ebook: Chemistry** Julia Burdge, 2014-10-16 Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems With outstanding art a consistent problem solving approach interesting applications woven throughout the chapters and a wide range of end of chapter problems this is a great third edition text

**Environmental Organic Chemistry** René P. Schwarzenbach, Philip M. Gschwend, Dieter M. Imboden, 2005-06-24 Environmental Organic Chemistry focuses on environmental factors that govern the processes that determine the fate of organic chemicals in natural and engineered systems The information discovered is then applied to quantitatively assessing the environmental behaviour of organic chemicals Now in its 2nd edition this book takes a more holistic view on physical chemical properties of organic compounds It includes new topics that address aspects of gas solid partitioning bioaccumulation and transformations in the atmosphere Structures chapters into basic and sophisticated sections Contains illustrative examples problems and case studies Examines the fundamental aspects of organic physical and inorganic chemistry applied to environmentally relevant problems Addresses problems and case studies in one volume

**Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

**Vadose Zone Processes** John S. Selker, James T. McCord, C. Kent Keller, 1999-06-28 Vadose Zone Processes provides a unified up to date treatment on the movement of water through unsaturated media In addition to covering the basic equations governing the flow and fate of water in unsaturated media the text covers the biogeochemistry of vadose environments and the statistical description of vadose processes The authors emphasize maintaining an intuitive understanding of how the results are derived and how they are appropriately applied This comprehensive and important book will be useful not only to those in traditional fields such as civil engineering geology crop science chemical engineering agricultural engineering and hydrology but also in the newer environmental engineering fields

including containment transport pollution remediation and waste disposal      Fundamentals of Atmospheric Modeling Mark Z. Jacobson, 1999 Comprehensive graduate text describing the atmospheric processes numerical methods and computational techniques needed for those studying air pollution and meteorology      **Solvent Extraction Principles and Practice, Revised and Expanded** Jan Rydberg, 2004-03-03 A complete and up to date presentation of the fundamental theoretical principles and many applications of solvent extraction this enhanced Solvent Extraction Principles and Practice Second Edition includes new coverage of the recent developments in solvent extraction processes the use of solvent extraction in analytical applications and waste re      Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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