

# Data and Calculations: The Atomic Spectrum of Hydrogen

## A. Calculation of the Energy Levels of the Hydrogen Atom

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

Table 11.2

| Quantum Number, $n$ | Energy, $E_n$ , in kJ/mole              | Quantum Number, $n$ | Energy, $E_n$ , in kJ/mole |
|---------------------|-----------------------------------------|---------------------|----------------------------|
| 1                   | $-2.16 \times 10^{-21}$                 | 6                   | $-6.01 \times 10^{-23}$    |
| 2                   | $-5.41 \times 10^{-19}$                 | 7                   | $4.42 \times 10^{-23}$     |
| 3                   | $-2.4 \times 10^{-22}$                  | 8                   | $2.67 \times 10^{-23}$     |
| 4                   | $-1.35 \times 10^{-22}$                 | 9                   | $2.67 \times 10^{-23}$     |
| 5                   | <del>8.00</del> $-8.66 \times 10^{-23}$ | 10                  | $2.16 \times 10^{-23}$     |

## B. Calculation of Wavelengths in the Spectrum of the H Atom

In the upper half of each box write  $\Delta E$ , the difference in energy in kJ/mole between  $E_{n_u}$  and  $E_{n_l}$ . In the lower half of the box, write  $\lambda$  in nm associated with that value of  $\Delta E$ .

Table 11.3

| $n_u$ | 6                                            | 5                                            | 4                                             | 3                                            | 2                                          | 1 |
|-------|----------------------------------------------|----------------------------------------------|-----------------------------------------------|----------------------------------------------|--------------------------------------------|---|
| 1     | $2.1 \times 10^{-21}$<br>$94.2 \text{ nm}$   | $2.07 \times 10^{-21}$<br>$95.6 \text{ nm}$  | $2.02 \times 10^{-21}$<br>$98.01 \text{ nm}$  | $1.92 \times 10^{-21}$<br>$105.1 \text{ nm}$ | $3.24 \times 10^{-22}$<br>$611 \text{ nm}$ |   |
| 2     | $4.8 \times 10^{-22}$<br>$412.5 \text{ nm}$  | $4.52 \times 10^{-22}$<br>$437.7 \text{ nm}$ | $4.05 \times 10^{-22}$<br>$488.8 \text{ nm}$  | $3 \times 10^{-22}$<br>$660 \text{ nm}$      |                                            |   |
| 3     | $1.8 \times 10^{-22}$<br>$1100 \text{ nm}$   | $1.5 \times 10^{-22}$<br>$1320 \text{ nm}$   | $1.05 \times 10^{-22}$<br>$1885.7 \text{ nm}$ |                                              |                                            |   |
| 4     | $7.5 \times 10^{-23}$<br>$2440 \text{ nm}$   | $4.8 \times 10^{-23}$<br>$4125 \text{ nm}$   |                                               |                                              |                                            |   |
| 5     | $2.6 \times 10^{-23}$<br>$7615.3 \text{ nm}$ |                                              |                                               |                                              |                                            |   |

$$\Delta E = E_{n_u} - E_{n_l}$$

$$\lambda (\text{nm}) = \frac{1.19627 \times 10^9}{\Delta E}$$

(continued on following page)

# Atomic Spectrum Of Hydrogen Lab Answers

**Brendan G. Carr**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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