

# Assigning Oxidation Numbers

| Rule                                                                                                                                      | Example                                                       |
|-------------------------------------------------------------------------------------------------------------------------------------------|---------------------------------------------------------------|
| 1. The oxidation number of any uncombined element is 0.                                                                                   | The oxidation number of $\text{Na}(s)$ is 0.                  |
| 2. The oxidation number of a monatomic ion equals the charge on the ion.                                                                  | The oxidation number of $\text{Cl}^-$ is -1.                  |
| 3. The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.      | The oxidation number of O in NO is -2.                        |
| 4. The oxidation number of fluorine in a compound is always -1.                                                                           | The oxidation number of F in LiF is -1.                       |
| 5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in a peroxide, in which it is -1. | The oxidation number of O in $\text{NO}_2$ is -2.             |
| 6. Hydrogen's oxidation state in most of its compounds is +1 unless it is combined with a metal, in which case it is -1.                  | The oxidation number of H in LiH is -1.                       |
| 7. In compounds, the elements of Groups 1 and 2 as well as aluminum have oxidation numbers of +1, +2, and +3, respectively.               | The oxidation number of Ca in $\text{CaCO}_3$ is +2.          |
| 8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.                                                              | The oxidation number of C in $\text{CaCO}_3$ is +4.           |
| 9. The sum of the oxidation numbers of all atoms in a polyatomic ion equals the charge of the ion.                                        | The oxidation number of P in $\text{H}_2\text{PO}_4^-$ is +5. |

# Assigning Oxidation Numbers Answers

**J Dewey**



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