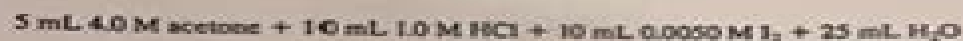


Experiment 20

Advance Study Assignment: The Iodination of Acetone

- In a reaction involving the iodination of acetone, the following volumes were used to make up the reaction mixture:



- How many moles of acetone were in the reaction mixture? Recall that, for a component A, moles $A = M_A \times V$, where M_A is the molarity of A and V is the volume in liters of the solution of A that was used. $5 \text{ mL} = 0.005 \text{ L}$

$$A = 4.0 \text{ M} \times 0.005 \text{ L}$$

$$\underline{0.02} \text{ moles acetone}$$

$$A = 0.02 \text{ moles acetone}$$

- What was the molarity of acetone in the reaction mixture? The volume of the mixture was 50 mL, 0.050 L, and the number of moles of acetone was found in Part (a). Again,

$$M_A = \frac{\text{moles of A}}{V \text{ of soln. in liters}}$$

$$M_A = \frac{0.02 \text{ M}}{0.050 \text{ L}} = 0.4 \text{ M acetone}$$

$$\underline{0.4} \text{ M acetone}$$

- How could you double the molarity of the acetone in the reaction mixture, keeping the total volume at 50 mL and keeping the same concentrations of H^+ ion and I_2 as in the original mixture?

1. Double the moles of acetone concentration

- Using the reaction mixture in Problem 1, a student found that it took 510 seconds for the color of the I_2 to disappear.

- What was the rate of the reaction? Hint: First find the initial concentration of I_2 in the reaction mixture, $[I_2]_0$. Then use Equation 3.

$$[I_2]_0 = 0.0050 \text{ M} \times 0.01 = \frac{0.000050 \text{ mol}}{0.050 \text{ L}} = 0.001 \text{ M}$$

$$\text{rate} = \underline{2.0 \times 10^{-6}} \text{ M/sec}$$

$$\text{rate} = \frac{0.001 \text{ M}}{510 \text{ s}} = 1.96 \times 10^{-6} \text{ M.s}^{-1} \approx 2.0 \times 10^{-6} \text{ M.s}^{-1}$$

- Given the rate from Part (a), and the initial concentrations of acetone, H^+ ion, and I_2 in the reaction mixture, write Equation 3 as it would apply to the mixture.

$$\text{rate} = K [\text{acetone}]^m [I_2]^n [H^+]^p$$

$$2.0 \times 10^{-6} = K (0.4 \text{ M})^m (0.02)^n (0.001)^p$$

- What are the unknowns that remain in the equation in Part (b)?

$$\underline{K} \quad \underline{m} \quad \underline{n} \quad \underline{p}$$

(continued on following page)

$$\begin{aligned} \text{mol } H^+ &= 1.0 \text{ M} \times 0.010 \\ &= \frac{0.010}{0.050} = 0.20 \end{aligned}$$

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