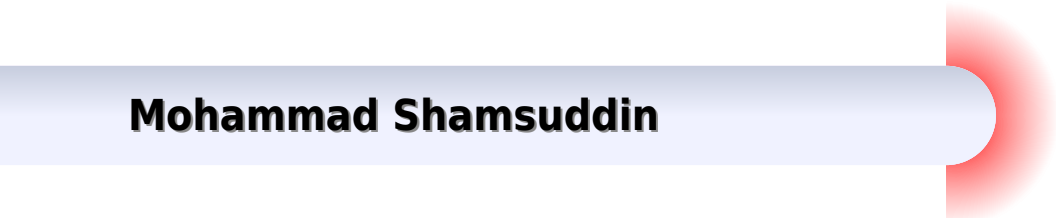


Some Basic Concept of Chemistry

1. Number of molecules in $W(g)$ of substance = $\frac{W(g) \times N_A}{GMM}$
2. Molality (m) = $\frac{\text{No. of moles of solute}}{\text{Mass of solvent in kg}}$
3. Number of molecules in V litre of gas at S.T.P. = $\frac{VN_A}{22.4}$
4. Number of gram atoms = $\frac{W(g)}{GAM}$ (GAM $\rightarrow$ gram atomic mass)
5. Number of gram molecules = $\frac{W(g)}{\text{Gram molecular mass}}$
6. Dilution formula: $M_1V_1 = M_2V_2$
For mixing two solutions of the same substance
 $M_1V_1 + M_2V_2 = M_3(V_1 + V_2)$
Molarity can be directly calculated from % by mass (w/w) if density is known
Molarity = $\frac{\% \times 10 \times d}{GMM}$
7. Mass of 1 atom of element = $\frac{GAM}{N_A}$
8. Mass of 1 molecule of substance = $\frac{MM}{N_A}$ (MM $\rightarrow$ Molar mass)
9. $T(K) = T(^{\circ}C) + 273.15$
10. Relative atomic mass = $\frac{\text{Mass of an atom of the element}}{\frac{1}{12} \times \text{Mass of an atom of carbon (C-12)}}$
11. Number of molecules in n moles of substance = $n \times N_A$
12. Mass % of an element in a compound = $\frac{\text{Mass of that element in 1 mole of the compound}}{\text{Molar mass of the compound}}$
13. Mass percent = $\frac{\text{Mass of solute}}{\text{Mass of solution}} \times 100$
14. $\frac{X_a}{1 - X_a} = \frac{\text{Molality} \times M_a}{1000}$ where M_a = mass of solvent
15. Molarity (M) = $\frac{\text{No. of moles of solute}}{\text{Volume of solution in litres}}$ mole / L.
16. Avogadro's No. $N_A = 6.022 \times 10^{23}$
17. $T(^{\circ}F) = \frac{9}{5} T(^{\circ}C) + 32$
18. Molecular mass = $2 \times \text{vapour density}$
19. Mole fraction of A = $\frac{\text{No. of moles of A}}{\text{No. of moles of solution}}$

Basic Physical Chemistry Second Edition

Mohammad Shamsuddin



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Basic Physical Chemistry for the Atmospheric Sciences Peter V. Hobbs, 2000-09-04 Revised and updated in 2000 Basic Physical Chemistry for the Atmospheric Sciences provides a clear concise grounding in the basic chemical principles required for studies of atmospheres oceans and earth and planetary systems Undergraduate and graduate students with little formal training in chemistry can work through the chapters and the numerous exercises within this book before accessing the standard texts in the atmospheric chemistry geochemistry and the environmental sciences The book covers the fundamental concepts of chemical equilibria chemical thermodynamics chemical kinetics solution chemistry acid and base chemistry oxidation reduction reactions and photochemistry In a companion volume entitled Introduction to Atmospheric Chemistry 2000 Cambridge University Press Peter Hobbs provides an introduction to atmospheric chemistry itself including its applications to air pollution acid rain the ozone hole and climate change Together these two books provide an ideal introduction to atmospheric chemistry for a variety of disciplines

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compact than standard textbooks on the subject and emphasises the two important topics underpinning the subject quantum mechanics and the second law of thermodynamics Both topics are challenging to students because they focus on uncertainty and probability The book explains these fundamental concepts clearly and shows how they offer the key to understanding a wide range of chemical phenomena including atomic and molecular spectra the structure and properties of solids liquids and gases chemical equilibrium and the rates of chemical reactions This revised edition has enabled improvements and corrections to be made Physical Chemistry of Metallurgical Processes, Second Edition Mohammad

Shamsuddin, 2021-06-22 This updated second edition retains its classroom tested treatment of physical chemistry of metallurgical topics such as roasting of sulfide minerals matte smelting converting structure properties and theories of slag reduction of oxides and reduction smelting interfacial phenomena steelmaking secondary steelmaking role of halides in extraction of metals refining hydrometallurgy and electrometallurgy and adds new data in worked out examples as well as up to date references to the literature The book further explains the physical chemistry of various metallurgical topics steps involved in extraction of metals such as roasting matte smelting converting reduction smelting steelmaking reactions deoxidation stainless steelmaking vacuum degassing refining leaching chemical precipitation ion exchange solvent extraction cementation gaseous reduction and electrowinning Each topic is illustrated with appropriate examples of applications of the technique in extraction of some common reactive rare or refractory metal together with worked out problems explaining the principle of the operation The problems require imagination and critical analyses and also encourage readers for creative application of thermodynamic data in metal extraction Updates and condenses text throughout the book by sequential arrangement of paragraphs in different chapters Maximizes readers understanding of the physicochemical principles involved in extraction production of common and rare reactive metals by pyro as well as hydrometallurgical routes Reinforces concepts presented with worked examples in each chapter explaining the process steps Explains the physical chemistry of various metallurgical steps such as roasting matte smelting converting and reduction smelting steelmaking aqueous processing etc in extraction of metals Collects and uniformly presents scattered information on physicochemical principles of metal production from various books and journals **Physical Properties of Materials, Second Edition**

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Principles of Physical Chemistry Hans Kuhn, Horst-Dieter Försterling, David H. Waldeck, 2009-03-17 *Principles of Physical Chemistry* Second Edition uniquely uses simple physical models as well as rigorous treatments for understanding molecular and supramolecular systems and processes In this way the presentation assists students in developing an intuitive understanding of the subjects as well as skill in quantitative manipulations The unifying nature of physical chemistry is emphasized in the book by its organization beginning with atoms and molecules and proceeding to molecular assemblies of increasing complexity ending with the emergence of matter that carries information i.e. the origin of life a physicochemical process of unique importance The aim is to show the broad scope and coherence of physical chemistry

Physical Chemistry William M. Davis, 2011-12-06 Designed specifically for a two semester introductory course sequence in physical chemistry this text presents core principles and topics Straightforward and streamlined it presents the necessary amount of detail for comprehension Organized in such a way that the various topics covered are connected to each other it allows students to see physical chemistry as an interconnected discipline and not a series of unrelated concepts Each chapter in this new edition has been thoroughly updated and includes new information on computational applications more end of chapter problems and new chapters on nanotechnology and surface chemistry

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 - The Rise of Digital Reading Basic Physical Chemistry Second Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic Physical Chemistry Second Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic Physical Chemistry Second Edition
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Physical Chemistry Second Edition
 - Personalized Recommendations
 - Basic Physical Chemistry Second Edition User Reviews and Ratings
 - Basic Physical Chemistry Second Edition and Bestseller Lists
5. Accessing Basic Physical Chemistry Second Edition Free and Paid eBooks
 - Basic Physical Chemistry Second Edition Public Domain eBooks
 - Basic Physical Chemistry Second Edition eBook Subscription Services
 - Basic Physical Chemistry Second Edition Budget-Friendly Options

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 - Basic Physical Chemistry Second Edition Enhanced eBook Features
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 - Highlighting and Note-Taking Basic Physical Chemistry Second Edition
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic Physical Chemistry Second Edition
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic Physical Chemistry Second Edition
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Basic Physical Chemistry Second Edition
 - Setting Reading Goals Basic Physical Chemistry Second Edition
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Basic Physical Chemistry Second Edition
 - Fact-Checking eBook Content of Basic Physical Chemistry Second Edition
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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