



Some basic concepts of chemistry



1 atomic mass unit (amu) = 1.66056×10^{-24}



percentage composition = $\frac{\text{Mass of that element in the compound}}{\text{Molar mass of compound}}$



Mass percent = $\frac{\text{Mass of solute}}{\text{Mass of solution}}$



Moles of an element = $\frac{\text{Mass of element}}{\text{Atomic mass}}$



Moles of one molecule = $\frac{\text{Molecular mass}}{6.022 \times 10^{23}}$



Molarity (M) = $\frac{\text{Moles of Solute}}{\text{Volume of solution in litre}}$



Molality (m) = $\frac{\text{Moles of Solute}}{\text{Mass of solution in kg}}$



Mole fraction of the solvent in the solution

$$x_a = \frac{n_a}{n_a + n_b}$$



Mole fraction of the solute in the solution

$$x_b = \frac{n_b}{n_a + n_b}$$



sum of the mole fraction of the component is unity

$$x_a + x_b = 1$$



Number of grams of atoms = $\frac{\text{Weight (in grams)}}{\text{gram atomic mass}}$



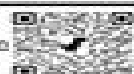
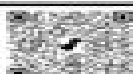
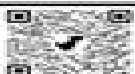
Gram molecular No. = $\frac{\text{Mass (in grams)}}{\text{Molecular weight}}$



Molecular Mass($A_x B_y C_z$) = $x[\text{Atomic mass of A}] + y[\text{Atomic mass of B}] + z[\text{Atomic mass of C}]$



Molecular Mass = $\frac{\text{Actual mass of a molecule of matter}}{1/12 \times \text{mass of one C-12 atom}}$



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