

3 Materialen

Eigenschappen en deeltjesmodellen | vwo

Newton

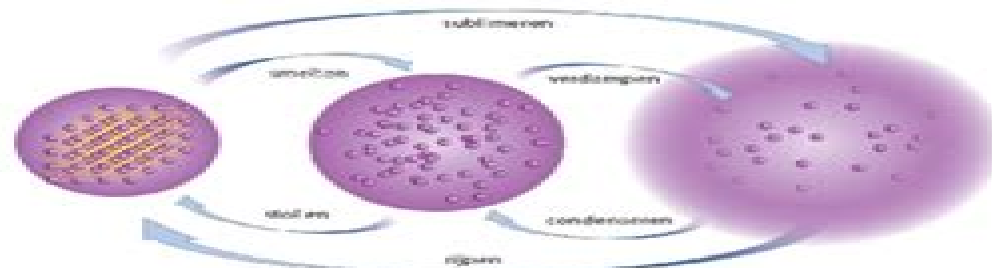
Uitwerkingen diagnostische toets



1

- a Vaste en vloeibare fase
- b Temperatuur is een maat voor de bewegingsenergie van de deeltjes van de stof. Bij $-273\text{ }^{\circ}\text{C}$ is de snelheid van deeltjes nul en is dus ook de bewegingsenergie nul. Langzamer bewegen dan niet bewegen kan niet.
- c Het onderscheid tussen een vaste en een vloeistof is dat deeltjes in een vaste stof op een vaste plek blijven, terwijl ze in een vloeistof langs elkaar kunnen bewegen. Bij 0 kelvin bewegen de deeltjes niet en kun je niet weten of de deeltjes langs elkaar heen kunnen bewegen.

2



Figuur 1

- a Beide deeltjes hebben dezelfde temperatuur en bewegen dus even snel. Een waterdeeltje beweegt daarbij door de stof heen terwijl het ijsdeeltje op één plek heen en weer trilt.
- b Bij opwarmen gaan de deeltjes steeds sneller bewegen. Als de stof verdampt dan beweegt een deeltje zo snel dat het ontsnapt aan de aantrekkingskracht van de rest van de deeltjes en de kriskras door de hele ruimte gaat bewegen. Bij koken gebeurt dit niet alleen aan het oppervlak van de vloeistof maar ook in het midden; hier botsen de deeltjes nu zo hevig hierdoor raken ze erg ver van elkaar verwijderd en trekken de deeltjes elkaar niet meer aan.
- d Beide deeltjes hebben dezelfde temperatuur en dus de dezelfde hoeveelheid bewegingsenergie. Het gouddeeltje is zwaarder dan het aluminiumdeeltje en beweegt langzamer bij dezelfde hoeveelheid bewegingsenergie.

3

- a De dichtheid wordt bepaald door de massa en het volume van het gas: $\rho = \frac{m}{V}$. De massa verandert niet, het volume wordt $2x$ zo klein. Dus wordt de dichtheid $2x$ zo groot.
- b Het volume verandert niet, terwijl het aantal deeltjes en dus de massa verdubbelt. De dichtheid wordt dus $2x$ zo groot.
- c Het volume van het gas blijft hetzelfde en er gaat ook geen deeltje de fles in of uit, dus blijft de massa gelijk. De dichtheid blijft hetzelfde.
- d Door de stijging van de temperatuur gaan de deeltjes meer bewegen en doordat ze harder botsen zullen ze zich op grote afstand van elkaar bevinden. Daardoor stroomt er gas uit de fles $\frac{p_1 \cdot V_1}{n_1 \cdot T_1} = \frac{p_2 \cdot V_2}{n_2 \cdot T_2}$; Hierbij veranderen de druk en het volume niet; $\frac{1 \times 1}{100\% \times 293} = \frac{1 \times 1}{n_2 \times 313}$.
 $n_2 = 94\%$. De massa is dus kleiner, het volume blijft gelijk. De dichtheid is dus 94% geworden.

4

- a $n = \frac{p \cdot V}{R \cdot T} = \frac{1,4 \cdot 10^5 \times 0,45}{8,31 \times 298} = 25 \text{ mol}$
- b $\frac{p_1 \cdot V_1}{n_1 \cdot T_1} = \frac{p_2 \cdot V_2}{n_2 \cdot T_2}$, invullen van de gegevens: $\frac{1,4 \times 0,45}{1 \times 298} = \frac{p_2 \times 0,52}{1 \times 291}$ dus: $p_2 = 1,2 \text{ bar}$

Antwoorden Diagnostische Toets Mens En Milieu

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