764	[A _j]	B ,	billal rate $(M \epsilon^{-1})$
1	0.10	0.50	2.5×10^{-6}
2	0.20	0.50	5.0 × 10 ⁻⁴
3	0.30	0.05	5.0 × 10 ⁻⁶
4	0.00	0.10	1.0×10^{-4}

An experiment was conducted to determine the rate law for the reaction $A_{x}(y) = B(y) - A_{x}B(y)$. The table above shows the data collected, Sased on the data in the table, which statement is correct?

- C Since the rate law can be expressed as $rate = E[A_{\omega}]^2[B]$, tripling the concentration of A_{ω} while keeping the concentration of B constant will triple the rate of the reaction.
- Since the rate law can be expressed as rate = k[A₂[B]², doubling the concentration of B while keeping the concentration of A₂ constant will double the rate of the reaction.

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