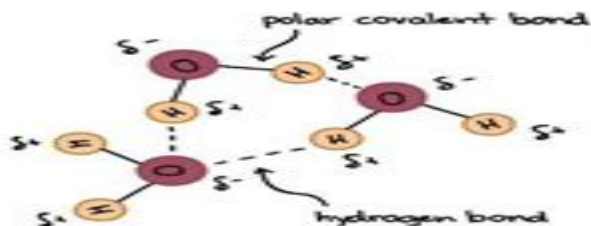


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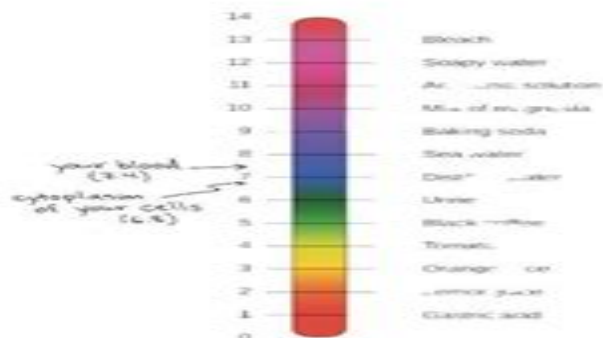
WATER



- The properties of water relate to its shape, polar nature and its ability to form **hydrogen bonds** with itself and with other molecules.
- Water has a **high specific heat capacity** and a **high heat of vaporisation**.
- Water molecules can ionise into H^+ and OH^- . **Ionisation** of water is important for living systems because of the reactive nature of H^+ .

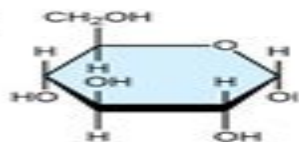
PH SCALE

- The pH scale is used to express degrees of acidity.
- An **acid** is a substance that causes the concentration of H^+ in a solution to rise above its equilibrium value.
- A **base** is a substance that causes a decrease in the relative numbers of H^+ and an increase in the relative numbers of OH^- in a solution.
- Buffers** are chemical systems that maintain a **relatively constant pH** even when substantial amounts of acid or base are added.

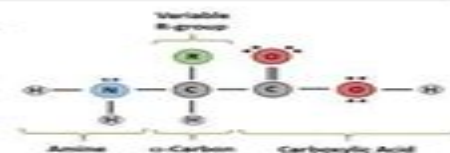


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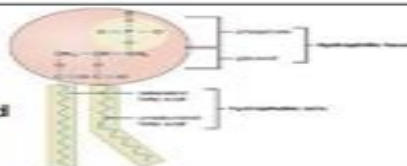
Carbohydrates are primarily a combination of carbon and water.



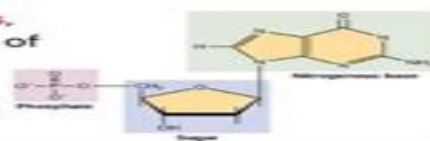
Proteins are polymers of amino acids arranged in linear sequence.



Lipids are molecules consisting of three fatty acid and glycerol.



Nucleic acids, are made out of units called nucleotides (DNA & RNA).



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 - The Rise of Digital Reading Biochemistry Study Guides
 - Advantages of eBooks Over Traditional Books
2. Identifying Biochemistry Study Guides
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Biochemistry Study Guides
 - User-Friendly Interface
4. Exploring eBook Recommendations from Biochemistry Study Guides
 - Personalized Recommendations
 - Biochemistry Study Guides User Reviews and Ratings
 - Biochemistry Study Guides and Bestseller Lists
5. Accessing Biochemistry Study Guides Free and Paid eBooks
 - Biochemistry Study Guides Public Domain eBooks
 - Biochemistry Study Guides eBook Subscription Services
 - Biochemistry Study Guides Budget-Friendly Options
6. Navigating Biochemistry Study Guides eBook Formats
 - ePub, PDF, MOBI, and More
 - Biochemistry Study Guides Compatibility with Devices
 - Biochemistry Study Guides Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Biochemistry Study Guides
 - Highlighting and Note-Taking Biochemistry Study Guides
 - Interactive Elements Biochemistry Study Guides
8. Staying Engaged with Biochemistry Study Guides

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Biochemistry Study Guides
- 9. Balancing eBooks and Physical Books Biochemistry Study Guides
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Biochemistry Study Guides
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Biochemistry Study Guides
 - Setting Reading Goals Biochemistry Study Guides
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Biochemistry Study Guides
 - Fact-Checking eBook Content of Biochemistry Study Guides
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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