

AKAI SERVICE MANUAL



MIDI PROGRAMMABLE PATCH BAY MODEL **ME30PII**

I. SPECIFICATIONS

Connection terminals		Indicator	red/green two colored LED × 8
INPUT	MIDI IN (DIN/5P) × 4	Memory bank	32
OUTPUT	MIDI THRU (DIN/5P) × 4	Power requirements	120V, 60Hz for USA and Canada 220V, 50Hz for Europe except UK 240V, 50Hz for UK and Australia
BANK UP/DOWN for foot switch/standard	1	Power consumption	40W
RCA	1	Front panel shape	ELA-1U rack mount
Function		Maximum external dimension	482.6 (W) × 45.7 (H) × 120 (D) mm
Power switch	1	Weight	2 kg
BANK UP switch	1		
BANK DOWN switch	1		
INPUT switch	1		
OUTPUT switch	8		
MIDI (No. 7 switch of OUTPUT)			
COPY (No. 8 switch of OUTPUT)			
Display	7 segmented red LED/two digit		

* For improvement purposes, specifications and design are subject to change without notice.

Akai Me30pii Manual

Brendan G. Carr



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web 1 practice problem balance the redox reaction in basic solution no 2 $\text{aq} + \text{S} + \text{NH}_3 \rightarrow \text{g} + \text{AlO}_2^{-} + \text{aq}$ 2 practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of OH^{-} br 2 l bro 3 aq br aq previous topic next topic

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web google classroom for the following redox equation $\text{MnO}_4^{-} + \text{aq} + \text{SO}_2 \rightarrow \text{g} + \text{Mn}^{2+} + \text{aq} + \text{HSO}_4^{-}$ a 4 a aq what will be the stoichiometric coefficient of HSO_4^{-} a 4 a if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^{-}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$ 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a

factor of 6

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