

Some Basic Concept of Chemistry

1. Number of molecules in $W(g)$ of substance = $\frac{W(g) \times N_A}{GMM}$
2. Molality (m) = $\frac{\text{No. of moles of solute}}{\text{Mass of solvent in kg}}$
3. Number of molecules in V litre of gas at S.T.P. = $\frac{VN_A}{22.4}$
4. Number of gram atoms = $\frac{W(g)}{GAM}$ (GAM $\rightarrow$ gram atomic mass)
5. Number of gram molecules = $\frac{W(g)}{\text{Gram molecular mass}}$
6. Dilution formula: $M_1V_1 = M_2V_2$
For mixing two solutions of the same substance
 $M_1V_1 + M_2V_2 = M_3(V_1 + V_2)$
Molarity can be directly calculated from % by mass (w/w) if density is known
Molarity = $\frac{\% \times 10 \times d}{GMM}$
7. Mass of 1 atom of element = $\frac{GAM}{N_A}$
8. Mass of 1 molecule of substance = $\frac{MM}{N_A}$ (MM $\rightarrow$ Molar mass)
9. $T(K) = T(^{\circ}C) + 273.15$
10. Relative atomic mass = $\frac{\text{Mass of an atom of the element}}{\frac{1}{12} \times \text{Mass of an atom of carbon (C-12)}}$
11. Number of molecules in n moles of substance = $n \times N_A$
12. Mass % of an element in a compound = $\frac{\text{Mass of that element in 1 mole of the compound}}{\text{Molar mass of the compound}}$
13. Mass percent = $\frac{\text{Mass of solute}}{\text{Mass of solution}} \times 100$
14. $\frac{X_a}{1 - X_a} = \frac{\text{Molality} \times M_A}{1000}$ where M_A = mass of solvent
15. Molarity (M) = $\frac{\text{No. of moles of solute}}{\text{Volume of solution in litres}}$ mole / L.
16. Avogadro's No. $N_A = 6.022 \times 10^{23}$
17. $T(^{\circ}F) = \frac{9}{5} T(^{\circ}C) + 32$
18. Molecular mass = $2 \times$ vapour density
19. Mole fraction of A = $\frac{\text{No. of moles of A}}{\text{No. of moles of solution}}$

Basic Physical Chemistry Second Edition

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