
AAMC MCAT Practice Exam 1 Questions and Answers

In the chromatography of the reaction mixture, water absorbed on cellulose functioned as the stationary phase. What was the principal factor determining the migration of individual components in the sample?

- A) Hydrogen bonding
 - B) Solute concentration
 - C) Stationary phase concentration
 - D) Thickness of paper
- Answer - A) Hydrogen bonding

The answer to this question is A because the relative amount of hydrogen bonding to the stationary phase will determine the relative rate of migration of the various components in the sample.

What assumption is being made if scientists conclude that aspartic acid was formed by the prebiological synthesis in the passage?

- A) Aspartic acid is unstable at temperatures below 150 degrees C.
 - B) All of the malic acid underwent the dehydration reaction to form fumaric/maleic acid.
 - C) Compound A and cyanide were available on primitive Earth.
 - D) The reaction between ammonia and fumaric acid was catalyzed by the presence of water.
- Answer - C) Compound A and cyanide were available on primitive Earth.

The answer to this question is C since, in order for the experimental reaction sequence to be relevant to the primordial formation of aspartic acid, the starting materials used (Compound A and cyanide) are assumed to have been available.

According to the developed chromatography plate shown below, what is the approximate R_f value of aspartic acid?

aspartic acid = 2
solvent front = 10

- A) 0.20
 - B) 0.50
 - C) 5
 - D) 10
- Answer - The answer to this question is A (0.20) because R_f is the ratio of the distance travelled by the analyte relative to the solvent front during a chromatographic separation. Aspartic acid travelled two units, while the solvent front travelled ten units, giving an R_f of $2/10 = 0.20$ for aspartic acid.

Which of the following statements does NOT correctly describe the

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Michael Brown



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1. Understanding the eBook Answers To Aamc Practice Test 3
 - The Rise of Digital Reading Answers To Aamc Practice Test 3
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Aamc Practice Test 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Aamc Practice Test 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Aamc Practice Test 3
 - Personalized Recommendations
 - Answers To Aamc Practice Test 3 User Reviews and Ratings

- Answers To Aamc Practice Test 3 and Bestseller Lists
- 5. Accessing Answers To Aamc Practice Test 3 Free and Paid eBooks
 - Answers To Aamc Practice Test 3 Public Domain eBooks
 - Answers To Aamc Practice Test 3 eBook Subscription Services
 - Answers To Aamc Practice Test 3 Budget-Friendly Options
- 6. Navigating Answers To Aamc Practice Test 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Aamc Practice Test 3 Compatibility with Devices
 - Answers To Aamc Practice Test 3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Aamc Practice Test 3
 - Highlighting and Note-Taking Answers To Aamc Practice Test 3
 - Interactive Elements Answers To Aamc Practice Test 3
- 8. Staying Engaged with Answers To Aamc Practice Test 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Aamc Practice Test 3
- 9. Balancing eBooks and Physical Books Answers To Aamc Practice Test 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Aamc Practice Test 3
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Aamc Practice Test 3
 - Setting Reading Goals Answers To Aamc Practice Test 3
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Aamc Practice Test 3
 - Fact-Checking eBook Content of Answers To Aamc Practice Test 3
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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