Assigning Oxidation Numbers

Rule	Example
 The oxidation number of any uncombined element is 0. 	The oxidation number of $Na(s)$ is 0.
The oxidation number of a monatomic ion equals the charge on the ion.	The oxidation number of Cl ⁻ is -1.
The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.	The oxidation number of O in NO is -2.
 The oxidation number of fluorine in a compound is always –1. 	The oxidation number of F in LiF is -1.
 Oxygen has an oxidation number of −2 unless it is combined with F, in which it is +1 or +2, or it is in a peroxide, in which it is −1. 	The oxidation number of O in NO ₂ is -2.
 Hydrogen's oxidation state in most of its compounds is +1 unless it is combined with a metal, in which case it is −1. 	The oxidation number of H in LiH is -1.
 In compounds, the elements of Groups 1 and 2 as well as aluminum have oxidation numbers of +1, +2, and +3, respectively. 	The oxidation number of Ca in CaCO ₃ is +2.
 The sum of the oxidation numbers of all atoms in a neutral compound is 0. 	The oxidation number of C in CaCO ₃ is +4.
The sum of the oxidation numbers of all atoms in a polyatomic ion equals the charge of the ion.	The oxidation number of P in H ₂ PO ₄ is +5.

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Dr Vijay Sarda, Dr A C Handa, Dr K K Arora

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