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Chapter 5: Atom and Ion Movements in Materials

- 5-10 Atoms are found to move from one lattice position to another at the rate of 5×10^5 jumps/s at 400°C when the activation energy for their movement is 30,000 cal/mol. Calculate the jump rate at 750°C .

Solution:

$$\begin{aligned}\text{Rate} &= \frac{5 \times 10^5}{x} = \frac{c \exp[-30,000/(1.987)/(673)]}{c_v \exp[-30,000/(1.987)/(1023)]} = \exp(-22.434 + 14.759) \\ \frac{5 \times 10^5}{x} &= \exp(-7.675) = 4.64 \times 10^{-4} \\ x &= \frac{5 \times 10^5}{4.64 \times 10^{-4}} = 1.08 \times 10^9 \text{ jumps/s}\end{aligned}$$

- 5-11 The number of vacancies in a material is related to temperature by an Arrhenius equation. If the fraction of lattice points containing vacancies is 8×10^{-5} at 600°C , determine the fraction of lattice points containing vacancies at 1000°C .

Solution: $8 \times 10^{-5} = \exp[-Q/(1.987)/(873)]$ $Q = 16,364 \text{ cal/mol}$
 $f = n_v/n = \exp[-16,364/(1.987)/(1273)] = 0.00155$

- 5-12 The Arrhenius equation was originally developed for comparing rates of chemical reactions. Compare the rates of a chemical reaction at 20 and 100°C by calculating the ratio of the chemical reaction rates. Assume that the activation energy for liquids in which the chemical reaction is conducted is 10 kJ/mol and that the reaction is limited by diffusion.

Solution:

An Arrhenius relationship gives the rate of a process or reaction as a function of temperature. In general, for an Arrhenius relationship,

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Jones, Ann Fretwell, Peter J. Goodhew**

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