

## CHAPTER 7

## STUDY GUIDE

## Ionic Compounds and Metals

### Section 7.1 Ion Formation

*In your textbook, read about chemical bonds and formation of ions.*

Use each of the terms below just once to complete the passage.

|               |           |                             |         |             |
|---------------|-----------|-----------------------------|---------|-------------|
| chemical bond | electrons | energy level                | ions    | noble gases |
| nucleus       | octet     | pseudo-noble gas formations | valence |             |

The force that holds two atoms together is called a(n) (1) \_\_\_\_\_.

Such an attachment may form by the attraction of the positively charged

(2) \_\_\_\_\_ of one atom for the negatively charged

(3) \_\_\_\_\_ of another atom, or by the attraction of charged atoms, which are called

(4) \_\_\_\_\_. The attractions may also involve

(5) \_\_\_\_\_ electrons, which are the electrons in the outermost

(6) \_\_\_\_\_. The (7) \_\_\_\_\_ are a family of elements that have very

little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a

stable (8) \_\_\_\_\_. The relatively stable electron structures developed by loss of

electrons in certain elements of groups 3, 4, 13, and 14 are called (9) \_\_\_\_\_.

For each statement below, write *true* or *false*.

- \_\_\_\_\_ 10. A positively charged ion is called an anion.
- \_\_\_\_\_ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- \_\_\_\_\_ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- \_\_\_\_\_ 13. A sodium atom tends to lose one electron when it reacts.
- \_\_\_\_\_ 14. The electron structure of a zinc ion ( $\text{Zn}^{2+}$ ) is an example of a pseudo-noble gas formation.
- \_\_\_\_\_ 15. A  $\text{Cl}^-$  ion is an example of a cation.
- \_\_\_\_\_ 16. The ending *-ide* is used to designate an anion.
- \_\_\_\_\_ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

# Answer Study Guide Ionic Compounds And Metals

**Christian G. Meyer**



## **Answer Study Guide Ionic Compounds And Metals:**

## **Answer Study Guide Ionic Compounds And Metals** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Answer Study Guide Ionic Compounds And Metals**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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