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CHAPTER 5: Electrons in Atoms Section 2: Quantum Theory and the Atom

Boler's Model of the Atom.

Energy states of hydrogen

He describes electrons in terms of their energy levels:

- Electrons can be only certain distances from the recleus.
- Each distance corresponds to a certain quantity of energy that an electron can have.
- An electron that is as close to the nucleus as it can is in its lowest energy level, the further an electron is from the nucleus, the higher the energy level that the electron occupies.
- The difference in energy between two energy levels is known as Quantum of energy.
- An electron can be in only one energy level or another not between energy levels.
- Normally if an electron is in a state of lowest possible energy, it is in a ground state.
- If an electron gains energy, it moves to an excited state.
- An electron in an excited state will release a specific quantity of energy as it quickly falls back to its ground state.
- The smaller the electron's orbit, the lower the atom's energy state, or energy level.
- Conversely, the larger the electron's orbit, the higher the atom's energy state, or energy level.
- Thus, a hydrogen atom can have many different excited states, although it contains only one electron.
- Bohr's idea is illustrated in figure 10

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