

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids:** Atoms, ions, or molecules are arranged in a regular, repeating three-dimensional pattern. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids:** Atoms, ions, or molecules are arranged in a disordered, non-repeating pattern. They do not have sharp melting points and do not show distinct X-ray diffraction patterns.
- **Intermolecular Forces:** The forces between particles in a solid determine its properties. Stronger forces lead to higher melting points and greater hardness.
- **Types of Solids:**
 - Atomic Solids:** Composed of atoms held together by metallic or covalent bonds.
 - Molecular Solids:** Composed of molecules held together by intermolecular forces.
 - Ionic Solids:** Composed of ions held together by ionic bonds.
 - Network Solids:** Composed of atoms held together by a continuous network of covalent bonds.

Types of Chemical Calculations Allowed on the Exam

- Stoichiometry
- Thermochemistry
- Equilibrium
- Kinetics
- Electrochemistry
- Acid-Base Chemistry
- Solubility
- Coordination Chemistry
- Organic Chemistry
- Analytical Chemistry
- Environmental Chemistry
- Materials Chemistry
- Biochemistry
- Polymer Chemistry
- Nanotechnology
- Environmental Chemistry
- Materials Chemistry
- Biochemistry
- Polymer Chemistry
- Nanotechnology

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy increases from bottom-left to top-right. Atomic radius increases from top-left to bottom-right.

- **Electronegativity:** The ability of an atom to attract electrons in a bond. It increases across a period and decreases down a group.
- **Ionization Energy:** The energy required to remove an electron from an atom. It increases across a period and decreases down a group.
- **Atomic Radius:** The distance from the nucleus to the outer edge of the electron cloud. It increases down a group and across a period.
- **Electron Affinity:** The energy change when an electron is added to a neutral atom. It increases across a period and decreases down a group.
- **Electron Configuration:** The arrangement of electrons in an atom's orbitals. It follows the Aufbau principle, Pauli exclusion principle, and Hund's rule.

Electronegativity Table

Electronegativity Table

Electronegativity values for various elements, categorized by groups.

Group	1	2	13	14	15	16	17	18
Period 1	Li: 1.0	Be: 1.5	B: 2.0	C: 2.5	N: 3.0	O: 3.5	F: 4.0	Ne: -
Period 2	Na: 0.9	Mg: 1.2	Al: 1.5	Si: 1.8	P: 2.1	S: 2.5	Cl: 3.0	Ar: -
Period 3	K: 0.8	Ca: 1.0	Ga: 1.6	Ge: 2.0	As: 2.2	Se: 2.4	Br: 2.8	Kr: -
Period 4	Rb: 0.7	Sr: 0.9	In: 1.7	Sn: 1.9	Sb: 2.1	Te: 2.3	I: 2.5	Xe: -
Period 5	Cs: 0.7	Ba: 0.9	Tl: 1.8	Pb: 2.1	Bi: 2.2	Po: 2.4	At: 2.6	Rn: -
Period 6	Fr: 0.7	Ra: 0.9	Pb: 2.1	Bi: 2.2	Po: 2.4	At: 2.6	At: 2.6	Rn: -
Period 7	Fr: 0.7	Ra: 0.9	Pb: 2.1	Bi: 2.2	Po: 2.4	At: 2.6	At: 2.6	Rn: -

Thermodynamics

First Law of Thermodynamics: Energy cannot be created or destroyed, only transferred or converted from one form to another.

Second Law of Thermodynamics: The total entropy of an isolated system can never decrease over time.

Third Law of Thermodynamics: The entropy of a perfect crystal at absolute zero is zero.

Enthalpy (H): A measure of the total energy of a system, including internal energy and work done by or on the system.

Entropy (S): A measure of the disorder or randomness of a system.

Gibbs Free Energy (G): A measure of the maximum work that can be done by a system at constant temperature and pressure.

Equilibrium & Solubility

Chemical Equilibrium: A state in which the concentrations of reactants and products remain constant over time.

Le Chatelier's Principle: If a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift to counteract the change.

K_c and K_p: Equilibrium constants for reactions in terms of concentration and partial pressure, respectively.

Q: The reaction quotient, used to determine the direction a reaction will proceed to reach equilibrium.

Reaction	K _c	K _p
H ₂ (g) + I ₂ (g) ⇌ 2HI(g)	50.0	50.0
N ₂ (g) + 3H ₂ (g) ⇌ 2NH ₃ (g)	0.10	0.10
PCl ₅ (g) ⇌ PCl ₃ (g) + Cl ₂ (g)	0.040	0.040
2SO ₂ (g) + O ₂ (g) ⇌ 2SO ₃ (g)	8.0	8.0
CO(g) + H ₂ O(g) ⇌ CO ₂ (g) + H ₂ (g)	1.0	1.0
2NO(g) + O ₂ (g) ⇌ 2NO ₂ (g)	4.6	4.6
2H ₂ (g) + O ₂ (g) ⇌ 2H ₂ O(g)	9.1	9.1
2H ₂ (g) + O ₂ (g) ⇌ 2H ₂ O(l)	1.5	1.5
2H ₂ (g) + O ₂ (g) ⇌ 2H ₂ O(s)	1.5	1.5

Solubility: The maximum amount of a substance that can dissolve in a given amount of solvent at a specific temperature and pressure.

K_{sp}: The solubility product constant, used to determine the solubility of a compound.

Q_{sp}: The ion product, used to determine the direction a reaction will proceed to reach equilibrium.



Acids & Bases: Substances that donate or accept protons (H⁺ ions).

Strong Acids: Acids that completely dissociate in water (e.g., HCl, HNO₃, H₂SO₄).

Weak Acids: Acids that partially dissociate in water (e.g., CH₃COOH, H₂CO₃).

Strong Bases: Bases that completely dissociate in water (e.g., NaOH, KOH, Ca(OH)₂).

Weak Bases: Bases that partially dissociate in water (e.g., NH₃, CH₃NH₂).

pH and pK_a: Measures of acidity and basicity.

Redox Reactions

Oxidation: The loss of electrons or an increase in oxidation state.

Reduction: The gain of electrons or a decrease in oxidation state.

Oxidation-Reduction (Redox) Reaction: A reaction in which one substance is oxidized and another is reduced.

Half-Reactions: The oxidation and reduction reactions occurring separately.

Electrochemistry: The study of the relationship between electrical energy and chemical reactions.

Galvanic Cell: A cell that converts chemical energy into electrical energy.

Standard Reduction Potentials (E°): A measure of the tendency of a half-reaction to occur as a reduction.

Nernst Equation: An equation that relates the cell potential to the concentrations of the reactants and products.

Faraday's Laws: Laws that relate the amount of substance produced or consumed in an electrochemical reaction to the amount of electricity used.

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Michael Seilmaier



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