

TESTS

EXPLORE CHEMISTRY WITH CHEMISTRY

DESCRIBING CHEMICAL REACTIONS

SOLUTIONS TO THE MODULE 7 REVIEW QUESTIONS

1. Elements cannot undergo decomposition reactions because elements have nothing to decompose into.
2. a. This equation represents a synthesis reaction because 2 elements come together to produce 1 compound.
b. This equation represents a formation reaction because 2 elements come together to produce 1 compound.
c. None of these.
d. None of these.
3. This equation represents a decomposition reaction because a single compound is breaking down into its elements.
4. The only equation that does not represent a formation reaction is (b). The equation for (b) is a formation equation because it starts with 2 reactants and makes only 1 in the end. Remember, reactions do not have to start with elements in order to be formation equations; we just assume they do when we write them ourselves.
5. The only equation that does not represent a combustion reaction is (b). The equation for (b) is a decomposition equation. Even though it produces CO_2 and H_2O , it does not have O_2 as a reactant, so it is not combustion.
6. Complete combustion produces carbon dioxide and water. Incomplete combustion produces carbon monoxide and water or carbon and water.

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