Experiment 25

Advance Study Assignment: pH, Its Measurement and Applications

t. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

Methyl violet Thymol blue

violet ≥1.6 Congo red orange≥2. Bromeresol green violet < 3.0 yellow < 4.0

Methyl yellow red < 4.0

What is the approximate pH of the solution?

215 Call ----

approx 31

- An aqueous solution of NH, has a pH of 11.6. The ammonia (NH,) molecule is the conjugate base of the NH₄" (called "anunonium") ion. Write the net ionic equation for the reaction that makes an aqueous solution of NH, basic. (Eq. 5.) NH₄ (aq) + OH (aq) -> NH₃ (aq) + H₂O(l)
- 3. The pH of a 0.10 M HOBr solution is 4.8.
 - n. What is [H1] in that solution?

1.58 × 10-5 M

b. What is [OBr 17 What is [HOBr]? (Where do the H* and OBr* ions come from?)

0-1015-1

_____M;______M

what is the value of K, for HOBr? What is the value of pK.?

Advance Study Assignment Experiment 25 Answers

Catherine A. Sanderson, Karen R. Huffman

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