Cheemistry 1C

Wilorksheet #3.

At 2200°C, K.= 0.050 for the reaction:

No fight
$$\bullet$$
 On fight \leftrightarrow 2 NO fight

a. What is the partial pressure of NO in equilibrium with N₂ and O₂ that was placed in a flusk at initial pressures of 0.80 and 0.20 atm?

	74.1	404	20000
Resident .	0.08	01.050	101
Change	1-00	100	+ 27%
Eigenblockers	0.85-30	0.20-2	2.0

Kp = (250)*/ ((0.80-x300.20-x01)

Solve by quadratic formula and get v= 0.009 atm or v=-0.052. So be you cannot have negative pressure, the answer is 0.009

- b. What happens to the concentration of the NO if we increase the temperature?
- c. What happens to the concentration of 0y when you add Ny at constant volume? Decreases
- d. What happens to the concentration of 0; when you add Ar at constant volume? No change
- A 1.00 L flask was filled with 2.0 mol SO₂ (g) and 2.0 mol NO₂ (g) and heated.
 After equilibrium was reached, it was found that 1.3 moles NO₂ was present.
 Assume that the reaction occurs under these conditions

$$SO_2(g) = NO_2(g) \Rightarrow SO_2(g) + NO(g)$$

a. Calculate the equilibrium constant. Kr. for the above reaction.

	501	NOv	500.	360
Printed	2.0	2.0	0	0
Change	198		+30	
Equilibrium	20-4	2.0-x	+ 10	+ 10

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Pei Kang Shen

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