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Chapter 1: Semiconductor Diodes

1. Copper has 29 orbiting electrons with only one electron in the outermost shell. The fact that the outermost shell with its 29th electron is incomplete (subshell can contain 2 electrons) and distant from the nucleus reveals that this electron is loosely bound to its parent atom. The application of an external electric field of the correct polarity can easily draw this loosely bound electron from its atomic structure for conduction.

Both intrinsic silicon and germanium have complete outer shells due to the sharing (covalent bonding) of electrons between atoms. Electrons that are part of a complete shell structure require increased levels of applied attractive forces to be removed from their parent atom.

2. Intrinsic material: an intrinsic semiconductor is one that has been refined to be as pure as physically possible. That is, one with the fewest possible number of impurities.

Negative temperature coefficient: materials with negative temperature coefficients have decreasing resistance levels as the temperature increases.

Covalent bonding: covalent bonding is the sharing of electrons between neighboring atoms to form complete outermost shells and a more stable lattice structure.

3.

4. a. $W = QV = (12 \mu C)(6 V) = 72 \mu J$

$$b. 72 \times 10^{-6} J = \frac{1 eV}{1.6 \times 10^{-19} J} = 2.625 \times 10^{14} eV$$

5. $48 eV = 48(1.6 \times 10^{-19} J) = 7.68 \times 10^{-18} J$
 $q = \frac{W}{V} = \frac{7.68 \times 10^{-18} J}{3.2 V} = 2.4 \times 10^{-19} C$

$6.4 \times 10^{-19} C$ is the charge associated with 4 electrons.

6.

GaP	Gallium Phosphide	$E_g = 2.24 eV$
ZnS	Zinc Sulfide	$E_g = 3.67 eV$

7. An *n*-type semiconductor material has an excess of electrons for conduction established by doping an intrinsic material with donor atoms having more valence electrons than needed to establish the covalent bonding. The majority carrier is the electron while the minority carrier is the hole.

A *p*-type semiconductor material is formed by doping an intrinsic material with acceptor atoms having an insufficient number of electrons in the valence shell to complete the

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Michael Brown



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