

# Structure, Bonding and properties



## 1. Why do atoms bond together?

Some atoms are very reluctant to combine with other atoms and exist in the air around us as single atoms. These are the **Noble Gases** and have **very stable electron arrangements** eg **2, 2.8** and **2.8.8** because their outer shells are full. The first three are shown in the diagrams below and explains why Noble Gases are so reluctant to form compounds with other elements.



**All other atoms therefore, bond together to become electronically more stable, that is to become like Noble Gases in electron arrangement. Bonding produces new substances and usually involves only the 'outer shell' or 'valency' electrons**

The phrase **CHEMICAL BOND** refers to the strong electrical force of attraction between the atoms or ions in the structure. The combining power of an atom is sometimes referred to as its **valency** and its value is linked to the number of outer electrons of the original uncombined atom.

**COVALENT BONDING** - **sharing electrons** to form **molecules** with **covalent bonds**, the bond is usually formed between **two non-metallic elements** in a molecule.

**IONIC BONDING** - By one atom **transferring electrons** to another atom.

**COORDINATE COVALENT BONDING** A coordinate covalent bond is special because it involves a shared pair of electrons that came from a single atom.

**METALLIC BONDING** The **crystal lattice of metals** consists of **ions NOT atoms** surrounded by a 'sea of electrons' forming **giant lattice**. These free or 'delocalised' electrons are the 'electronic glue' holding the particles together. There is a **strong electrical force of attraction** between these **mobile electrons (-)** and the 'immobile' **positive metal ions (+)** and this is the **metallic bond**.

- An ion is an atom or group of atoms carrying an overall **positive** or **negative** charge
  - eg  $\text{Na}^+$ ,  $\text{Cl}^-$ ,  $[\text{Cu}(\text{H}_2\text{O})]^{2+}$ ,  $\text{SO}_4^{2-}$  etc.
- If a particle, as in a neutral atom, has equal numbers of protons (+) and electrons (-) the particle charge is zero ie no overall electric charge.
- The proton number in an atom does not change BUT the number of associated electrons can!
- If negative electrons are removed the excess charge from the protons produces an overall positive ion.
- If negative electrons are gained, there is an excess of negative charge, so a negative ion is formed.
- The charge on the ion is numerically related to the number of electrons transferred.
- For any atom or group of atoms, for every electron gained you get a one unit increase in negative charge, for every electron lost you get a one unit increase in the positive

The atom **losing electrons** forms a **positive ion (cation)** and is usually a **metal**.

The atom **gaining electrons** forms a **negative ion (anion)** and is usually a **non-metallic element**.

# Atoms And Bonding Assessment Study Guide

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### **Table of Contents Atoms And Bonding Assessment Study Guide**

1. Understanding the eBook Atoms And Bonding Assessment Study Guide
  - The Rise of Digital Reading Atoms And Bonding Assessment Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Atoms And Bonding Assessment Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atoms And Bonding Assessment Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atoms And Bonding Assessment Study Guide
  - Personalized Recommendations
  - Atoms And Bonding Assessment Study Guide User Reviews and Ratings
  - Atoms And Bonding Assessment Study Guide and Bestseller Lists

5. Accessing Atoms And Bonding Assessment Study Guide Free and Paid eBooks
  - Atoms And Bonding Assessment Study Guide Public Domain eBooks
  - Atoms And Bonding Assessment Study Guide eBook Subscription Services
  - Atoms And Bonding Assessment Study Guide Budget-Friendly Options
6. Navigating Atoms And Bonding Assessment Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Atoms And Bonding Assessment Study Guide Compatibility with Devices
  - Atoms And Bonding Assessment Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Atoms And Bonding Assessment Study Guide
  - Highlighting and Note-Taking Atoms And Bonding Assessment Study Guide
  - Interactive Elements Atoms And Bonding Assessment Study Guide
8. Staying Engaged with Atoms And Bonding Assessment Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Atoms And Bonding Assessment Study Guide
9. Balancing eBooks and Physical Books Atoms And Bonding Assessment Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Atoms And Bonding Assessment Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Atoms And Bonding Assessment Study Guide
  - Setting Reading Goals Atoms And Bonding Assessment Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atoms And Bonding Assessment Study Guide
  - Fact-Checking eBook Content of Atoms And Bonding Assessment Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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