

REDOX REACTIONS.

NOTES...

Intro

~Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$
- Sub-points:
 - Loss of hydrogen
 - Gain of oxygen
 - Gain of **electronegative** element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$
- Sub-points:
 - Gain of hydrogen
 - Loss of oxygen
 - Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

~REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$

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G Psacharopoulos



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Table of Contents Balancing Redox Equations Study Guide Pg 117

1. Understanding the eBook Balancing Redox Equations Study Guide Pg 117
 - The Rise of Digital Reading Balancing Redox Equations Study Guide Pg 117
 - Advantages of eBooks Over Traditional Books
2. Identifying Balancing Redox Equations Study Guide Pg 117
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Balancing Redox Equations Study Guide Pg 117
 - User-Friendly Interface
4. Exploring eBook Recommendations from Balancing Redox Equations Study Guide Pg 117
 - Personalized Recommendations
 - Balancing Redox Equations Study Guide Pg 117 User Reviews and Ratings
 - Balancing Redox Equations Study Guide Pg 117 and Bestseller Lists
5. Accessing Balancing Redox Equations Study Guide Pg 117 Free and Paid eBooks
 - Balancing Redox Equations Study Guide Pg 117 Public Domain eBooks
 - Balancing Redox Equations Study Guide Pg 117 eBook Subscription Services
 - Balancing Redox Equations Study Guide Pg 117 Budget-Friendly Options
6. Navigating Balancing Redox Equations Study Guide Pg 117 eBook Formats
 - ePub, PDF, MOBI, and More
 - Balancing Redox Equations Study Guide Pg 117 Compatibility with Devices
 - Balancing Redox Equations Study Guide Pg 117 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Balancing Redox Equations Study Guide Pg 117
 - Highlighting and Note-Taking Balancing Redox Equations Study Guide Pg 117
 - Interactive Elements Balancing Redox Equations Study Guide Pg 117

8. Staying Engaged with Balancing Redox Equations Study Guide Pg 117
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Balancing Redox Equations Study Guide Pg 117
9. Balancing eBooks and Physical Books Balancing Redox Equations Study Guide Pg 117
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Balancing Redox Equations Study Guide Pg 117
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Balancing Redox Equations Study Guide Pg 117
 - Setting Reading Goals Balancing Redox Equations Study Guide Pg 117
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Balancing Redox Equations Study Guide Pg 117
 - Fact-Checking eBook Content of Balancing Redox Equations Study Guide Pg 117
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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