

Practice problems on gas laws

Practice problems on Boyle's law

Practice 1: calculate the final pressure of gas at the specified conditions assuming the temperature and mass remain constant. $V_1 = 1575 \text{ cm}^3$, $V_2 = 1050 \text{ cm}^3$ and $P_1 = 1.0 \text{ atm}$.

Practice 2: A 10.0 L volume of He gas exerts a pressure of 0.75 atm. What pressure is exerted by this sample of He if the volume occupied is decreased to 8.0 L at constant temperature?

Practice 3: In the following, calculate the final pressure of gas at the specified conditions assuming the temperature and mass remain constant. (a) $V_1 = 1575 \text{ cm}^3$, $V_2 = 1050 \text{ cm}^3$ and $P_1 = 1.0 \text{ atm}$. (b) $V_1 = 48 \text{ litres}$, $V_2 = 32 \text{ litres}$ and $P_1 = 1.8 \text{ atm}$. (c) $V_1 = 22 \text{ m}^3$, $V_2 = 40 \text{ m}^3$ and $P_1 = 80 \text{ kPa}$.

Practice 4: A fixed mass of gas at a constant temperature has a pressure of 2000 Pa and a volume of 0.02 m^3 . It is compressed until the volume is 0.005 m^3 . What is its new pressure?

Practice 5: Initially gas has a volume of 0.14 m^3 and a pressure of 300 kPa. What will be its volume when the pressure becomes 60 kPa if the temperature and mass remain unchanged?

Practice 6: A gas with a volume of 2 m^3 is compressed from a pressure of 100 kPa to a pressure of 500 kPa. If the temperature remains unchanged, what is the resulting volume?

Problems on Charles' law

Practice 1: Calculate the initial/final temperature of gas at the following specified conditions assuming that pressure and mass remain constant. (a) $V_1 = 30 \text{ litres}$, $V_2 = 25 \text{ litres}$ and $T_1 = 39^\circ\text{C}$. (b) $V_1 = 56 \text{ dm}^3$, $V_2 = 40 \text{ dm}^3$ and $T_2 = -13^\circ\text{C}$.

Practice 2: Calculate the initial volume of gas at the following specified conditions assuming that pressure and mass remain constant. (a) $V_2 = 100 \text{ cm}^3$, $T_1 = 270 \text{ K}$ and $T_2 = 300 \text{ K}$. (b) $V_2 = 25 \text{ dm}^3$, $T_1 = 240 \text{ K}$ and $T_2 = 330 \text{ K}$.

Practice 3: The volume of a fixed mass of gas at 27°C is 20 cm^3 . Find its volume at 57°C if the pressure is kept constant.

Practice 4: At a temperature of 200 K, the volume of a sample of gas in a cylinder with a smoothly fitting piston is 0.0024 m^3 . The cylinder is heated while allowing the gas to expand at constant pressure. Calculate the volume of the gas at a temperature of 300 K.

Practice 5: A gas has a volume of 0.010 m^3 at 18°C . What will be its volume at 85°C if the pressure acting on the gas remains unchanged?

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