Half-Lives

- 1. Radioactive materials undergo nuclear decay at a CONSTONY rate.
- 2. The half-life is the time required for na 4 of a sample to radioactive decay.
- 3. The half-life of lodine 131 is about 8 days. This means in eight days, 1314 of a sample would Permater.
 - a. Start with 200 mg of Iodine 131. How many remains after eight days? 100 Mg

How many days is this? 100% -> 50% -> 75% -> 12.5%

c. How long would it take for 87.5% of sodine to decay?



Half-Life Formula:

A= Aa (%)*

 $\alpha = t/h$

- 1. Label the variables above. $A_0 = initial \ \partial pacts + h = total time passing h = time for one half-like$ n = number of half-lives
- Nobelium 259 has a half-life of 58 minutes. How much remains from a 1000g sample after 1440 minutes?

 $h = \frac{1}{h} = \frac{1990}{59} = 29.0$ $A = 1000(\frac{1}{2})^{29.8}$ A = 1000($\frac{1}{2}$) A = 105 g remaining

Dubnium - 269 has a half-life of 34 seconds. How long will you wait to see 500g of it decay to 1g?

109 (0,002)= 109 (615) 7 (+=305 s A = A (1) h 」= 500(2) = 0,002(元) = →-2.7 = -0,3(元)

Practice Problems using the Formula (SHOW YOUR WORK)

The half-life of polonium - 218 is 3 min. If you start with 485g of polonium, how much remains after 22 minutes?

A= 405 (1) = 405 (2) 7.3 = 3,000 left

The half-life of polonium - 218 is 3 min. How much time will it take to see 25g decay to 3g?

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