

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before 2 mol 0 mol 0 mol

Change -2 mol +2 mol

After 0 mol 2 mol

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before 9.94 mol xs mol 0 mol

Change -9.94 mol +9.94 mol

After 0 mol 9.94 mol

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before 1.0 mol xs mol 0 mol 0 mol

Change -1.0 mol -6.0 mol

After 0 mol xs mol

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

Answer Key Chemical Equations And Stoichiometry

Mike Jess



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Table of Contents Answer Key Chemical Equations And Stoichiometry

1. Understanding the eBook Answer Key Chemical Equations And Stoichiometry
 - The Rise of Digital Reading Answer Key Chemical Equations And Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemical Equations And Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chemical Equations And Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemical Equations And Stoichiometry
 - Personalized Recommendations
 - Answer Key Chemical Equations And Stoichiometry User Reviews and Ratings
 - Answer Key Chemical Equations And Stoichiometry and Bestseller Lists
5. Accessing Answer Key Chemical Equations And Stoichiometry Free and Paid eBooks
 - Answer Key Chemical Equations And Stoichiometry Public Domain eBooks
 - Answer Key Chemical Equations And Stoichiometry eBook Subscription Services
 - Answer Key Chemical Equations And Stoichiometry Budget-Friendly Options
6. Navigating Answer Key Chemical Equations And Stoichiometry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Chemical Equations And Stoichiometry Compatibility with Devices
 - Answer Key Chemical Equations And Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Chemical Equations And Stoichiometry
 - Highlighting and Note-Taking Answer Key Chemical Equations And Stoichiometry
 - Interactive Elements Answer Key Chemical Equations And Stoichiometry
8. Staying Engaged with Answer Key Chemical Equations And Stoichiometry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Key Chemical Equations And Stoichiometry
- 9. Balancing eBooks and Physical Books Answer Key Chemical Equations And Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chemical Equations And Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Chemical Equations And Stoichiometry
 - Setting Reading Goals Answer Key Chemical Equations And Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Chemical Equations And Stoichiometry
 - Fact-Checking eBook Content of Answer Key Chemical Equations And Stoichiometry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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