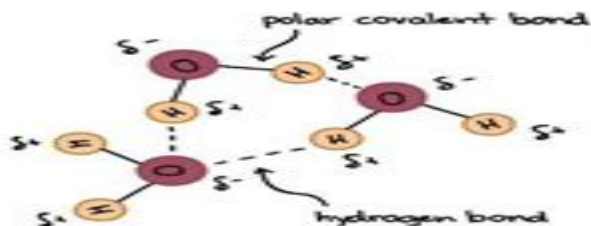


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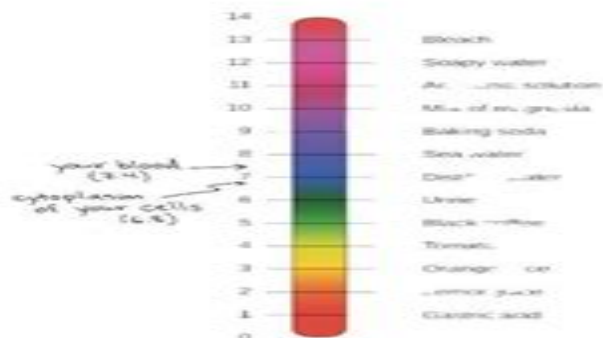
WATER



- The properties of water relate to its shape, polar nature and its ability to form **hydrogen bonds** with itself and with other molecules.
- Water has a **high specific heat capacity** and a **high heat of vaporisation**.
- Water molecules can ionise into H^+ and OH^- . **Ionisation** of water is important for living systems because of the reactive nature of H^+ .

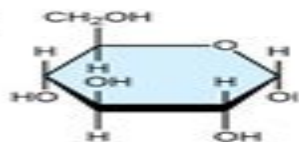
PH SCALE

- The pH scale is used to express degrees of acidity.
- An **acid** is a substance that causes the concentration of H^+ in a solution to rise above its equilibrium value.
- A **base** is a substance that causes a decrease in the relative numbers of H^+ and an increase in the relative numbers of OH^- in a solution.
- Buffers** are chemical systems that maintain a **relatively constant pH** even when substantial amounts of acid or base are added.

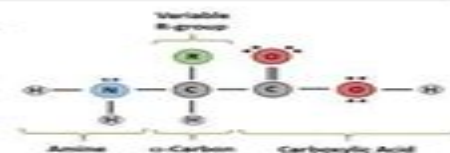


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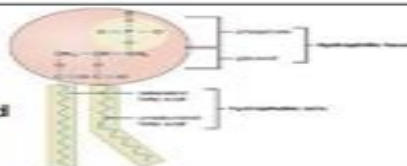
Carbohydrates are primarily a combination of carbon and water.



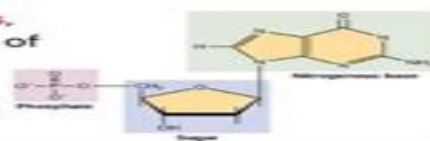
Proteins are polymers of amino acids arranged in linear sequence.



Lipids are molecules consisting of three fatty acid and glycerol.



Nucleic acids, are made out of units called nucleotides (DNA & RNA).



Biochemistry Study Guides

**Marcia L. Gillette, Ira Blei, Wendy
Gloffke, George Odian**



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