

CHEMISTRY ADVANCED STUDY

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FUNDAMENTAL CONCEPTS

Math

- Properties of exponents
- Properties of logarithms
 - $\log_b (M^p N) = \log_b (M) + \log_b (N)$
 - $\log_b (MN) = \log_b (M) + \log_b (N)$
 - $\log_b (M^p) = p \log_b (M)$
- Graphing slope intercepts
 - $y = mx + b$
 - b is the y-intercept, m is the slope
- Quadratic formula
 - $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$
 - Ex. $x^2 + 4x - 21$
 - $x = \frac{-4 \pm \sqrt{16 - 4(1)(-21)}}{2}$
 - $x = -2 \pm 5$
 - $x = 3$ or $x = -7$

Science

- Scientific notation
- Dimensional analysis
- Significant figures

Physics and Biology

- Scientific method
- Coulomb's law: electric force between like or opposite charges; dependence on separation distance and charge magnitude

ELEMENTS AND ATOMS

Protons: define the element (atomic number), positive

Electrons: define the charge (bonds and polarity), negative

Neutrons: define isotopes, neutral

ATOMIC MASS AND WEIGHT

Atomic Mass: Carbon-12 (exactly 12 atomic mass unit)

Atomic Weight

- weighted average of atomic masses of element
- (atomic mass · neutrons) + (atomic mass · neutrons)

MOLE A.K.A. AVOGADRO'S NUMBER

$$= 6.02 \cdot 10^{23}$$

In Terms of Carbon

1 mole of ^{12}C atoms in 12g of Carbon

$6.02 \cdot 10^{23} = ^{12}\text{C}$ atoms in 12g of Carbon

Gram to AMU

1 gram = 1 mole of atomic mass units

1 gram = $6.02 \cdot 10^{23}$ amu

1 mol ^{27}Al = 27g

2 mol ^{28}Si = 56g

4 mol ^{16}O = 64g

Atoms And Moles Study Guide

Leo J. Malone, Theodore O. Dolter



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Industrial Hygiene Reference & Study Guide Allan K. Fleeger, Dean R. Lillquist, 2006 *The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry** Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text **Chemical Principles Study Guide/Solutions Manual** John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement "O" Level Study Guide - Chemistry Quite Easily Done , **Study**

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 - Chapter 4: Atoms And Moles Study Guide in Specific Contexts
 - Chapter 5: Conclusion
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