



Atoms and molecules

an atom and its structure

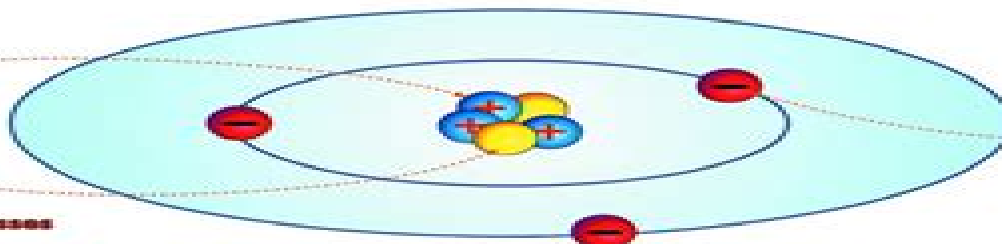
Atoms are made up of particles of matter known as **subatomic** or **fundamental particles**. The three main subatomic particles are called **protons**, **neutrons** and **electrons**. An atom consists mostly of empty space. Almost all of the atom's mass is concentrated in a structure at its centre called the **nucleus**. The nucleus consists of protons and neutrons grouped together. Electrons orbit the nucleus.

A **proton** is a positively charged particle

A **neutron** has no electric charge

Protons and neutrons have almost the same masses

An atom usually contains an equal number of positively charged protons and negatively charged electrons. This makes it electrically neutral.



An **electron** is a negatively charged particle. It is kept in orbit around the nucleus because it is attracted to the positively charged protons.

This attraction holds the atom together. An electron's mass is approximately 1/1840 times smaller than the mass of a proton or neutron.

electrons and their energy levels

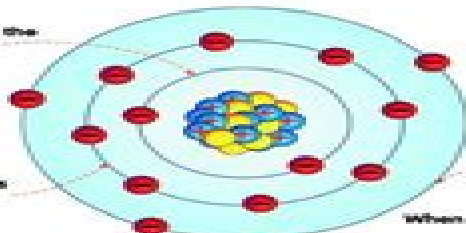
The way in which an atom reacts with other atoms (its **chemical properties**) is determined by the configuration of its electrons. Atoms can gain, lose or share electrons when they combine to form molecules. The electrons in an atom exist in layers called **energy levels** or **shells**. Electrons are arranged in the energy levels according to certain rules:

The first shell outside the nucleus holds up to two electrons

The second shell holds up to eight electrons

The third shell holds up to eight electrons.

When a shell is full, a new one is started.



Atomic Number	13
Atomic Symbol	Al
Atomic Mass	26.981 538

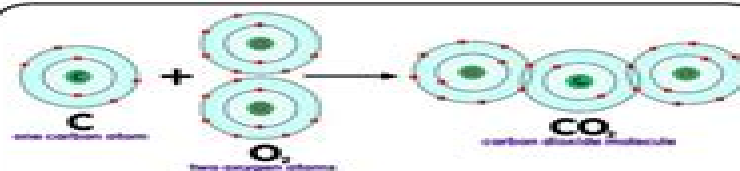
Number of electrons per shell (electron configuration)

The **atomic number** of an element tells you how many electrons that element has. An atom has as many energy levels as it needs to arrange its electrons. The number of electrons in the outer energy level determines how reactive an atom is. Atoms like their energy levels to be full. The most stable atoms are those with full outer energy levels, like helium. The most reactive atoms, like sodium, are those that only need to lose or gain one electron to obtain a full outer shell.

molecules

While an atom is the smallest particle of an element, **molecules** are the smallest particles of a compound. A molecule is formed when at least two atoms combine. These may be of the same element, or of different elements. If the atoms in a molecule are of different elements, it is called a compound.

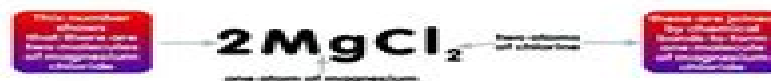
An example of such a compound is the carbon dioxide molecule.



A carbon atom has 4 electrons on its outer shell, while an oxygen atom has 6 electrons on its outer shell. It is thus necessary for 2 oxygen atoms to bond with 1 carbon atom so that the 2 oxygen atoms can fill their outer shells with 2 electrons, respectively, from the carbon atom ensuring that both oxygen atoms have full outer shells (8 required for 2nd shell) as does the carbon atom (2 required for 1st shell).

compounds & formulae

Compounds are formed when different elements combine. Like elements, each compound has a name and a symbol. The symbol of a compound is made of the symbols of the elements which form it, and is called a **formula**. A formula is used to show the atoms in a molecule. The numbers in a formula give the number of atoms of each kind of element in the molecule. The number of atoms of each element in that molecule is written after the symbol of the element, in small letters below the line e.g. O^2 which is a molecule of oxygen. A number written immediately before the formula of a molecule tells you how many molecules there are.



differences between mixtures & compounds

MIXTURES (e.g. iron and sulphur)

- Impure substances
- The proportions of the parts can vary (changed)
- The components can be elements or compounds
- The components are loosely put together
- The properties of each part do not change
- The properties of the mixture are the same as the properties of all its parts.

COMPOUNDS (e.g. iron sulphide)

- Pure substances
- The proportions of the parts are constant
- The components are elements
- The components are held together by chemical bonds.
- The properties of the compound are different from the properties of the elements forming it, and from a mixture of these elements.

The mixture formed when one substance **dissolves** into another is called a **solution**. Sugar that is dissolved in water is a solution. There are always two parts of a solution: the **solute** and the **solvent**. The component which forms the greatest part of the solution is the solvent. Water is the **solvent** of the sugar solution. The **sugar** is the **solute**.

Atoms And Molecules A Childs Guide To Chemistry

Clemens Wendtner



Atoms And Molecules A Childs Guide To Chemistry:

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Resources in Education, 1998

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