

Half-Lives

1. Radioactive materials undergo nuclear decay at a constant rate.
2. The half-life is the time required for half of a sample to radioactive decay.
3. The half-life of Iodine - 131 is about 8 days. This means in eight days, half of a sample would remain.

a. Start with 200 mg of Iodine - 131. How many remains after eight days? 100 mg

$$200 \text{ mg} \xrightarrow[\text{HL}]{8 \text{ d}} 100 \text{ mg}$$

b. You have 50 mg of Iodine - 131 remaining. How many half-lives have passed? 2

How many days is this? 16

$$100\% \rightarrow 50\% \rightarrow 25\% \rightarrow 12.5\%$$

c. How long would it take for 87.5% of Iodine to decay?

$$3(\text{HL}) = 24 \text{ days}$$

Half-Life Formula:

$$A = A_0 \left(\frac{1}{2} \right)^n$$

$$n = t/h$$

1. Label the variables above.
 A = amount left
 A_0 = initial amount
 n = number of half-lives
 t = total time passing
 h = time for one half-life
2. Nobelium - 259 has a half-life of 58 minutes. How much remains from a 1000g sample after 1440 minutes?

$$n = \frac{t}{h} = \frac{1440}{58} = 24.8 \quad A = 1000 \left(\frac{1}{2} \right)^{24.8} \rightarrow 3.36 \times 10^{-5} \text{ g remaining}$$

3. Dubnium - 269 has a half-life of 34 seconds. How long will you wait to see 500g of it decay to 1g?

$$A = A_0 \left(\frac{1}{2} \right)^n \quad \log(0.002) = \log(0.5)^{\frac{t}{34}} \quad t = 305 \text{ s}$$

$$\frac{1}{500} = \frac{500 \left(\frac{1}{2} \right)^{\frac{t}{34}}}{500} \rightarrow 0.002 \left(\frac{1}{2} \right)^{\frac{t}{34}} \rightarrow \frac{-2.7}{-0.3} = \frac{-0.3 \left(\frac{1}{2} \right)^{\frac{t}{34}}}{-0.3} \rightarrow (34) 8.97 = \left(\frac{t}{34} \right)^{34}$$

Practice Problems using the Formula (SHOW YOUR WORK)

1. The half-life of polonium - 218 is 3 min. If you start with 485g of polonium, how much remains after 22 minutes?

$$A = 485 \left(\frac{1}{2} \right)^{\frac{22}{3}} = 485 \left(\frac{1}{2} \right)^{7.3} = 3.08 \text{ g left}$$

2. The half-life of polonium - 218 is 3 min. How much time will it take to see 25g decay to 3g?

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