



# Batavus Course Manual

**JR Anderson**



## **Batavus Course Manual:**

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v_m \text{ dfrac } 32 \text{ g mol } 1 \text{ 790 g l } 1 \text{ 0 0405 l mol } 1$  nonumber

web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contain  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as  $6 \times 10^{23}$  particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always  $6 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web mar 12 2018 you can define the concentration of a gas though and that calculation would be exactly the one you've described placing  $x$  moles of a gas in a container of volume  $v$  yields a concentration  $c = x/v$  with units of  $\text{g mol}^{-1} \text{dm}^{-3}$

web the molar volume has the si unit of cubic metres per mole  $\text{m}^3 \text{mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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