

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HAc}} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-6}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-6} = 6 - \lg 2,8 = 6 - 0,45 = 5,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HAc}}}$$

$$= \sqrt{7,9 \cdot 10^{-6} \cdot 10^{-2}} = 2,8 \cdot 10^{-6} \text{ моль/л}$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-} = 2,5 \cdot 10^{-3}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}][\text{OH}^-] = 2,5 \cdot 10^{-3} \cdot 2,5 \cdot 10^{-3} = 6,25 \cdot 10^{-6}$$

S - концентрация ионов/л

$$S = \sqrt[3]{\frac{K_{\text{Pr}}}{4}} = \sqrt[3]{\frac{6,25 \cdot 10^{-6}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[OH^-] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 12 - \lg 1,72 = 11,76$$

③ Дано:

$$V_{\text{HAc}} = 10 \text{ мл}$$

$$C_{\text{HAc}} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 10 \text{ мл}$$

$$C_{\text{PrOH}_2} = 0,1 \text{ м}$$

$$K_{\text{PrHCl}_2} = 1,6 \cdot 10^{-5}$$

Решение:

полн. смешивание:

$$C_{\text{HAc}} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

Для гидролизующей соли:

$$C_{\text{H}^+} = C_{\text{HAc}} = 0,05 \text{ моль/л}$$

$$C_{\text{Pr}^{2+}} = C_{\text{PrOH}_2} = 0,05 \text{ моль/л}$$

$$\text{PrHCl}_2 \rightleftharpoons \text{Pr}^{2+} + 2\text{Cl}^-$$

$$K_{\text{Pr}} = [\text{Pr}^{2+}][\text{Cl}^-]^2 = 4 \cdot 0,05 \cdot 0,05 = 1 \cdot 10^{-5}$$

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Mary Pride



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