AP Chemistry: Acid-Base Worksheet

AP Chemistry Acid-Base WS 0809.doc

Name: ______ Date: _____ Per: ____

 In the following equations, label each species as an acid or a base. Show the conjugate acid-base pairs.

a. $HCO_3^-(eq) + HF_{(eq)} \leftrightarrow H_2CO_{3(eq)} + F_{(eq)}$ b. $HCO_3^-(eq) + OH^-_{(eq)} \leftrightarrow CO_3^{-2}_-(eq) + H_3O_{(f)}$

2. In the following reactions, identify the Lewis acid and the Lewis base.

a. $Ag^* + 2NH_2 \leftrightarrow Ag(NH_3)_2$

b. B(OH)₃ + H₂O ↔ B(OH)₄ + H*

For the following reaction, decide which species (reactants or products) are favored at the completion of the reaction.

 $SO_{1}^{2}(aq) + HCN(aq) \leftrightarrow HSO_{1}(aq) + CN'(aq)$

4. Which member of each of the following pairs is the stronger acid?

a. NH_a, PH_a

e. HSO, , HSeO,

b. HI, H.Te

c. HSO₃ , H₃SO₃

d. H-AsO., H-AsO.

5. Calculate the concentrations of hydronium ion and hydroxide ion at 25°C in

a. 0.15M HNO,

b. 0.010M Ca(OH),

- 6. A sample of orange juice has a hydronium-ion concentration of 2.9x10⁻¹M. What is the pH? Is the solution acidic?
- 7. The pH of human arterial blood is 7.40. What is the hydronium ion concentration?

Ap Worksheet 06d Acids And Bases Answer

Albert S. Tarendash, Frederick J. Rowe

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