

$$\begin{aligned}
 B' &= \frac{B}{RT} = \frac{-(110 \text{ cm}^3 \text{ mol}^{-1})(10^{-3} \text{ L cm}^{-3})}{(0.08314 \text{ L bar K}^{-1} \text{ mol}^{-1})(400 \text{ K})} \\
 &= -3.31 \times 10^{-3} \text{ bar}^{-1} \\
 Z &= 1 + B'P = 1 - (3.31 \times 10^{-3} \text{ bar}^{-1})(50 \text{ bar}) \\
 &= 0.835
 \end{aligned}$$

7. 1.5 Show that the virial equation written in terms of pressure has the following form for a van der Waals gas  $Z = 1 + [b - (a/RT)](P/RT)$  if terms in  $P^2$ ,  $P^3$ , etc., are neglected.

SOLUTION

$$P = \frac{RT}{V-b} - \frac{a}{V^2} \qquad \frac{PV}{RT} = \frac{V}{V-b} - \frac{a}{VRT}$$

Adding and subtracting  $\frac{V-b}{V-b}$  on the right side of the equation

$$\frac{PV}{RT} = 1 + \frac{b}{V-b} - \frac{a}{VRT}$$

Since the second and third terms are small correction terms,  $b$  can be ignored in comparison with  $V$ ; and  $V$  in the second and third term can be replaced by  $RT/P$ .

$$Z = 1 + \frac{bP}{RT} - \frac{aP}{(RT)^2} = 1 + \left[ b - \frac{a}{RT} \right] \left[ \frac{P}{RT} \right]$$

- 1.6 Show that for a gas of spherical molecules  $b$  in the van der Waals equation is four times the molecular volume times Avogadro's constant.

SOLUTION

The molecular volume for a spherical molecule is

$$\frac{4}{3}\pi\left(\frac{d}{2}\right)^3 = \frac{\pi}{6}d^3$$

# Alberty Physical Chemistry Manual Solution

**Robert A. Alberty**



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### **Table of Contents Alberty Physical Chemistry Manual Solution**

1. Understanding the eBook Alberty Physical Chemistry Manual Solution
  - The Rise of Digital Reading Alberty Physical Chemistry Manual Solution
  - Advantages of eBooks Over Traditional Books
2. Identifying Alberty Physical Chemistry Manual Solution
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Alberty Physical Chemistry Manual Solution
  - User-Friendly Interface
4. Exploring eBook Recommendations from Alberty Physical Chemistry Manual Solution
  - Personalized Recommendations
  - Alberty Physical Chemistry Manual Solution User Reviews and Ratings
  - Alberty Physical Chemistry Manual Solution and Bestseller Lists

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  - Alberty Physical Chemistry Manual Solution Enhanced eBook Features
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  - Highlighting and Note-Taking Alberty Physical Chemistry Manual Solution
  - Interactive Elements Alberty Physical Chemistry Manual Solution
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  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Alberty Physical Chemistry Manual Solution
9. Balancing eBooks and Physical Books Alberty Physical Chemistry Manual Solution
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Alberty Physical Chemistry Manual Solution
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Alberty Physical Chemistry Manual Solution
  - Setting Reading Goals Alberty Physical Chemistry Manual Solution
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Alberty Physical Chemistry Manual Solution
  - Fact-Checking eBook Content of Alberty Physical Chemistry Manual Solution
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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