Data and Calculations: The Atomic Spectrum of Hydrogen

A. Calculation of the Energy Levels of the Hydrogen Atom

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

Table 11.2							
Quantum Number, st	Energy, E, in klemole	Quantum Number, n	Energy, E_a , in kJ/mole				
1	-2.16 x 10-21	6	-6.01 VIO-33				
2	-5,41 × 10-17	7	4.45 XIO. 3				
3	-2,4 x10-22	8	2.67 x10-23				
ч	-1.35×10-22	9	2 67 X10-33				
5	-8.PF X10_72	10	2.16 × 10.23				

B. Calculation of Wavelengths in the Spectrum of the H Atom

In the upper half of each box write ΔE , the difference in energy in kJ/mole between E_{∞} and E_{∞} . In the lower half of the box, write λ in nm associated with that value of ΔE .

Table 11	Table 11.3						
Mad	6	5	4	3	2	1	
1	2.1 ×10 %	2.07 yie	2.02 10 1	1.92 x10 m	3.24vio		
-	94.2 00	95.6 nm	98,0100	103. Inm	611 nm		
2		4.53×10 4					
		437.7 nm					
3	1.8 y10 - 32	1.5x0-2	1.05x10				
	1100 nm	1320 nm	1885.7 An				
4	7.5 XIO-3	4.8 XID-23					
	240nm	4125nm					
	2.641033			Δ.E =	En - En		
	7615.3 m			λ (nm) =	1.19627 × 10 ΔΕ		

Atomic Spectrum Of Hydrogen Lab Answers

Ian Pickup

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